Date

# Chemistry — Unit 5 Worksheet 3 Empirical and Molecular Formulas

Show all your work when solving the following problems. Be sure to include units and label your answer.

 Find the empirical formula of a compound containing 32.0 g of bromine and 4.9 g of magnesium.

32.0 g × <u>1 mole</u> = 0.401 mole Br 79.9 g

1:2 ratio

4.9 g × 1 mole = 0.20 mole Mg 24.31 g

MgBrz

What is the empirical formula of a carbon-oxygen compound, given that a 95.2 g sample of the compound contains 40.8 g of carbon and the rest oxygen?
95.2 g - 40.8 g = 54.4 g O

 $40.8 \text{ g} \times 1 \text{ mole} = 3.40 \text{ moles C}$ 12.0 g

1:1 ratio

54.4 g x× <u>1 mole</u> = 3.40 moles O 16.0g CO

3. A compound was analyzed and found to contain 9.8 g of nitrogen, 0.70 g of hydrogen, and 33.6 g of oxygen. What is the empirical formula of the compound?

9.8 g × 1 mole = 0.70 moles N 14.0g

1:1:3 ratio

0.70 g × <u>1 mole</u> = 0.69 moles H 1.01 g HNO<sub>3</sub>

33.6 g × 1 mole = 2.10 moles O 16.0 g

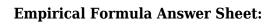
 A compound composed of hydrogen and oxygen is found to contain 0.59 g of hydrogen and 9.40 g of oxygen. The molar mass of this compound is 34.0 g/mol. Find the empirical and molecular formulas.

0.59 g × 1 mole = 0.58 mole H 1.01 g 9.40 g × 1 mole = 0.588 mole O 1:1 ratio; empirical formula: HO empirical formula mass: 17.0 g/mol molecular formula mass: 34.0 g/mol mfm is 2x efm so the molecular formula is: H-O-

16.0 g

# **Empirical Formula Answer Sheet**

**K Morrison** 



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## **Table of Contents Empirical Formula Answer Sheet**

- 1. Understanding the eBook Empirical Formula Answer Sheet
  - The Rise of Digital Reading Empirical Formula Answer Sheet
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Empirical Formula Answer Sheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Empirical Formula Answer Sheet
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Empirical Formula Answer Sheet

- Personalized Recommendations
- Empirical Formula Answer Sheet User Reviews and Ratings
- Empirical Formula Answer Sheet and Bestseller Lists
- 5. Accessing Empirical Formula Answer Sheet Free and Paid eBooks
  - Empirical Formula Answer Sheet Public Domain eBooks
  - Empirical Formula Answer Sheet eBook Subscription Services
  - Empirical Formula Answer Sheet Budget-Friendly Options
- 6. Navigating Empirical Formula Answer Sheet eBook Formats
  - o ePub, PDF, MOBI, and More
  - Empirical Formula Answer Sheet Compatibility with Devices
  - Empirical Formula Answer Sheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Empirical Formula Answer Sheet
  - Highlighting and Note-Taking Empirical Formula Answer Sheet
  - Interactive Elements Empirical Formula Answer Sheet
- 8. Staying Engaged with Empirical Formula Answer Sheet
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Empirical Formula Answer Sheet
- 9. Balancing eBooks and Physical Books Empirical Formula Answer Sheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Empirical Formula Answer Sheet
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Empirical Formula Answer Sheet
  - Setting Reading Goals Empirical Formula Answer Sheet
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Empirical Formula Answer Sheet

- Fact-Checking eBook Content of Empirical Formula Answer Sheet
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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