

Chemistry Honors Worksheet: Formulas of a Compound

1.
 - a.) moles The subscripts in a chemical formula represent the ratio of (moles / mass) in a compound
 - b.) CO Which compound has the greatest percentage of carbon? (CO, CO₂, PbCO₃)
 - c.) CH₄ What is the empirical formula of a compound that contains 0.2 moles of C and 0.8 moles of H
 - d.) 83 g/mol Calculate the molecular mass of a compound if 0.02461 moles has a mass of 2.05 g.
 - e.) Same The % composition of an empirical formula is (more / less / the same) than/as its molecular formula.
 - f.) one How many moles of V are indicated in one mole of the compound V(C₂O₄)₂?
 - g.) 1×10^{24} atoms A beaker of water contains 2×10^{24} atoms of H, how many atoms of O are also in the beaker?
 - h.) 6.02×10^{23} 1 mole of a molecular compound has a mass of 165 grams. How many molecules are there?
 - i.) 0% Fool's Gold is iron pyrite, FeS. What is the percentage of gold in a 1.0 gram sample of fool's gold?

2. The pesticide DDT has been banned in the United States since 1962. This compound has a formula C₁₄H₉Cl₅. Answer all questions referring to DDT. Note: Use 35.5 g/mol for Cl

- a.) molecular Is this compound ionic or molecular?
- b.) 5.2×10^{15} How many molecules of DDT are in a 6.1×10^{-4} g sample?
- c.) 1.53×10^{17} atoms H How many atoms of H are in 2.00 mg of DDT?
- d.) 7.51 g Cl How many grams of Cl are contained in 10.0 grams of DDT?
- e.) 7.84×10^{21} atoms C If a sample 5.04×10^{21} atoms of H, how many atoms of C would there also be?
- f.) 2.97×10^{20} atoms C How many atoms of C are in a 25.0 mg of DDT?

3. A compound of tantalum, Ta, and oxygen was produced in a lab by heating tantalum in a crucible. The data was collected.

Mass of crucible:	24.22 g		
Mass of crucible and tantalum:	26.14 g	> 1.92 g Ta	0.0106 mol Ta
Mass of crucible and tantalum oxide:	26.56 g	> 0.42 g O	0.02625 mol O

- a.) Calculate the percent composition of the compound. 82.1% Ta 17.9% O
- b.) What is the empirical formula of this compound? TaO_{2.5} - Ta₂O₅
- c.) Is this compound ionic or covalent? ionic
- d.) Name this compound: Tantalum(V) oxide * Need Roman Numerals!!!

4. Determine the empirical formula of the following compounds:

- a.) 63.0g Rb, 5.90g O Rb₂O
- b.) 4.95mg Th, 1.37mg S ThS₂
- c.) 2.13g Na, 2.32g As, 1.98g O Na₃AsO₄
- d.) 32.8% Cr, 67.2% Cl CrCl₃
- e.) 58.0% Rb, 9.50% N, 32.5% O RbNO₃

Empirical Molecular Formula Answers

JA Banks



Empirical Molecular Formula Answers:

Understanding General Chemistry Atef Korchef, 2022-03-07 Understanding General Chemistry details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter. Written in an easy to understand format with helpful pedagogy to fuel learning, the book features main objectives at the beginning of each chapter, get smart sections, and a check your reading section at the end of each chapter. The text is filled with examples and practices that illustrate the concepts at hand. In addition, a summary and extensive MCQs, exercises, and problems with the corresponding answers and explanations are readily available. Additional features include Alerts students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes. Offers answers and comments alongside sample problems enabling students to self-evaluate their skill level. Includes powerful methods, easy steps, simple and accurate interpretations, and engaging applications to help students understand complex principles. Provides a bridge to more complex topics such as solid state chemistry, organometallic chemistry, chemistry of main group elements, inorganic chemistry, and physical chemistry. This introductory textbook is ideal for chemistry courses for non-science majors as well as health sciences and preparatory engineering students. [E3 Chemistry Review Book - 2018 Home Edition \(Answer Key Included\)](#) Effiong Eyo, 2017-10-20 With Answer Key to All Questions. Chemistry students and homeschoolers: Go beyond just passing. Enhance your understanding of chemistry and get higher marks on homework, quizzes, tests, and the regents exam with E3 Chemistry Review Book 2018. With E3 Chemistry Review Book, students will get clear, clear, engaging, exciting, and easy to understand high school chemistry concepts with emphasis on New York State Regents Chemistry. The Physical Setting. Easy to read format to help students easily remember key and must-know chemistry materials. Several example problems with solutions to study and follow. Several practice multiple choice and short answer questions at the end of each lesson to test understanding of the materials. 12 topics of Regents question sets and 3 most recent Regents exams to practice and prep for any Regents Exam. This is the Home Edition of the book. Also available in School Edition. ISBN 978 197836229. The Home Edition contains an answer key section. Teachers who want to recommend our Review Book to their students should recommend the Home Edition. Students and parents whose school is not using the Review Book as instructional material as well as homeschoolers should buy the Home Edition. The School Edition does not have an answer key in the book. A separate answer key booklet is provided to teachers with a class order of the book. Whether you are using the school or Home Edition, our E3 Chemistry Review Book makes a great supplemental instructional and test prep resource that can be used from the beginning to the end of the school year. PLEASE NOTE: Although reading contents in both the school and home editions are identical, there are slight differences in question numbers, choices, and pages between the two editions. Students whose school is using the Review Book as instructional material SHOULD NOT buy the Home Edition. Also available in paperback print. **E3 Chemistry Guided Study Book - 2018 Home Edition (Answer Key Included)** Effiong

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Counting Moles Nigel P. Freestone, 2016-01-04 Students studying chemistry often struggle with the mole Counting Moles provides an effective aid to learning by giving clear and confident presentation of the essentials of the mole concept needed by those starting chemistry courses This user friendly self teach e book is split into six chapters which sequentially introduce the mole calculating frame to help solve problems Over 200 fully worked examples are given along with several hundred questions The mole concept is applied to topics such as relative atomic mass and relative formula mass percentage composition empirical and molecular formula The book also covers concentration its units volumetric analysis and the relationship between volume mass and moles of gases Counting Moles culminates in you taking a Mole Driving Test On passing this test you are issued with a Counting Moles Driving License that will give you all the confidence required to correctly answer all mole calculations

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