

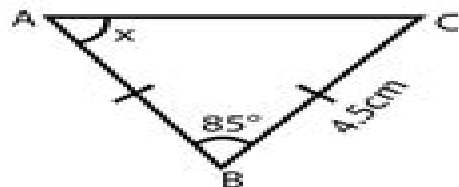
Name : _____

Score : _____ Date : _____

Isosceles Triangle Worksheet

Find the missing parameter(s) from the following isosceles triangle

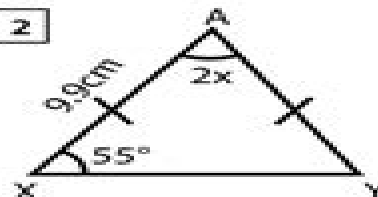
1



$x =$ _____

$AB =$ _____

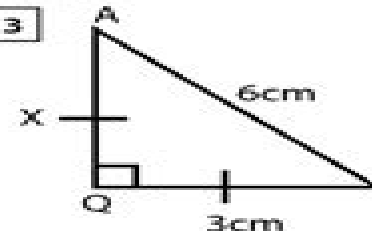
2



$x =$ _____

$AY =$ _____

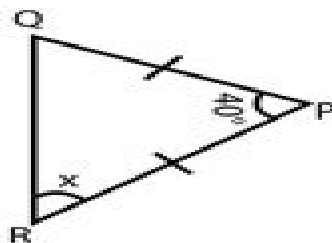
3



Perimeter = _____

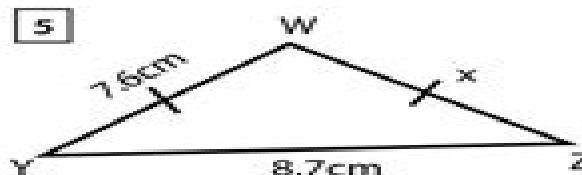
$x =$ _____

4



$x =$ _____

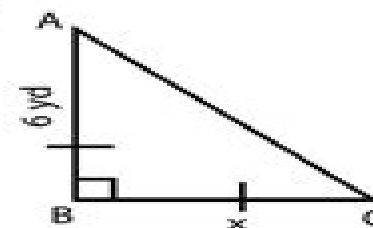
5



$x =$ _____

Perimeter = _____

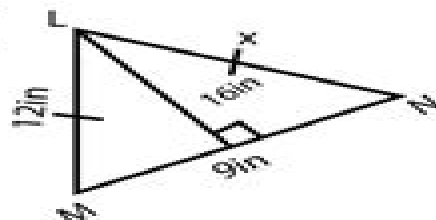
6



$x =$ _____

Area = _____

7

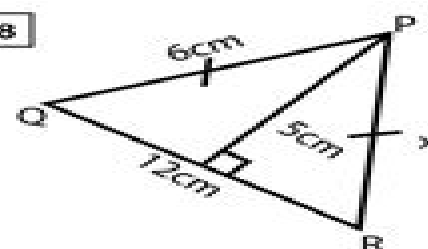


$x =$ _____

Area = _____

Perimeter = _____

8



$x =$ _____

Area = _____

Perimeter = _____

Mathematics Geometry Unit 05 Right Triangles Quiz

K Morrison



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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-}$ NO_3^- SO_4^{2-} MnO_2 Cr^{3+} NO_3^- SO_2 Mn_2O_3 balance each redox reaction in acid solution using the half reaction method 8 H_2O_2 $\text{Cr}_2\text{O}_7^{2-}$ 2 9 TeO_3 2 N_2O_4 O_2 Cr^{3+} Te 10 ReO_4^- 11 PbO_2 IO_3^- I^- Pb^{2+} 12 AsO_3^{3-} IO_3^- Re IO_3^- H_2AsO_4^-

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H^+ H_2O H^- and e^- in this order balance the atoms in the equation apart from O and H

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ $2\text{H}_2\text{NO}_2$ $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}_2$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2^- a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2^- b NO_2^- and $\text{C}_6\text{H}_8\text{O}_6$ c H_2O_2 and $\text{C}_6\text{H}_8\text{O}_6$ d H_2O_2 and NO_2^-