

## EMPIRICAL AND MOLECULAR FORMULA

- 1) What is the empirical formula of a compound that contains 46.3 % lithium and 53.7% oxygen?
- 2) What is the empirical formula of a compound that contains 15.9 % boron and 84.1 % fluorine?
- 3) Phosphorus reacts with oxygen to give a compound that is 43.7% phosphorus and 56.4% oxygen. What is the empirical formula of the compound?
- 4) An inorganic salt is composed of 17.6% sodium, 39.7% chromium and 42.8% oxygen. What is the empirical formula of this salt?
- 5) Compound X contains 69.9% carbon, 6.86% hydrogen and 23.3% oxygen. Determine the empirical formula of X.
- 6) Oxalic acid has the empirical formula  $\text{CHO}_2$ . Its molar mass is 90 g/mol. What is the molecular formula of oxalic acid?
- 7) The empirical formula of codeine is  $\text{C}_{18}\text{H}_{21}\text{NO}_3$ . If the molar mass of codeine is 299 g/mol, what is its molecular formula?
- 8) A compound's molar mass is 240.28 g/mol. Its percentage composition is 75.0% carbon, 5.05% hydrogen and 20.0% oxygen. What is the compound's molecular formula?
- 9) The wintergreen plant produces methyl salicylate, or oil of wintergreen. It can also be prepared easily in the laboratory. Methyl salicylate is 63.1% carbon, 5.31% hydrogen and 31.6% oxygen. Calculate the empirical formula of this compound.
- 10) An inorganic salt is made up of 38.8% calcium, 20.0% phosphorus, and 41.2% oxygen.
  - a) What is the empirical formula of the compound?
  - b) On further analysis, each formula unit of this salt is found to contain two phosphate ions. Predict the molecular formula of this salt.

### Answers

- 1)  $\text{Li}_2\text{O}$
- 2)  $\text{BF}_3$
- 3)  $\text{P}_2\text{O}_5$
- 4)  $\text{Na}_2\text{Cr}_2\text{O}_7$
- 5)  $\text{C}_{12}\text{H}_{14}\text{O}_3$
- 6)  $\text{C}_2\text{H}_2\text{O}_4$
- 7)  $\text{C}_{18}\text{H}_{21}\text{NO}_3$
- 8)  $\text{C}_{15}\text{H}_{12}\text{O}_3$
- 9)  $\text{C}_8\text{H}_8\text{O}_3$
- 10) a)  $\text{Ca}_3\text{P}_2\text{O}_8$  b)  $\text{Ca}_3(\text{PO}_4)_2$

# Empirical And Molecular Formula Problems

**LP Steffe**



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## **Table of Contents Empirical And Molecular Formula Problems**

1. Understanding the eBook Empirical And Molecular Formula Problems
  - The Rise of Digital Reading Empirical And Molecular Formula Problems
  - Advantages of eBooks Over Traditional Books
2. Identifying Empirical And Molecular Formula Problems
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Empirical And Molecular Formula Problems
  - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical And Molecular Formula Problems
  - Personalized Recommendations
  - Empirical And Molecular Formula Problems User Reviews and Ratings
  - Empirical And Molecular Formula Problems and Bestseller Lists
5. Accessing Empirical And Molecular Formula Problems Free and Paid eBooks
  - Empirical And Molecular Formula Problems Public Domain eBooks
  - Empirical And Molecular Formula Problems eBook Subscription Services
  - Empirical And Molecular Formula Problems Budget-Friendly Options

6. Navigating Empirical And Molecular Formula Problems eBook Formats
  - ePub, PDF, MOBI, and More
  - Empirical And Molecular Formula Problems Compatibility with Devices
  - Empirical And Molecular Formula Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Empirical And Molecular Formula Problems
  - Highlighting and Note-Taking Empirical And Molecular Formula Problems
  - Interactive Elements Empirical And Molecular Formula Problems
8. Staying Engaged with Empirical And Molecular Formula Problems
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Empirical And Molecular Formula Problems
9. Balancing eBooks and Physical Books Empirical And Molecular Formula Problems
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Empirical And Molecular Formula Problems
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Empirical And Molecular Formula Problems
  - Setting Reading Goals Empirical And Molecular Formula Problems
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Empirical And Molecular Formula Problems
  - Fact-Checking eBook Content of Empirical And Molecular Formula Problems
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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