

## CP Chemistry Unit 1 Test Review KEY

### Problems and questions:

- Round each of the following numbers to 3 significant figures:  
a) .09882 **.0988**      b) 123,459 **123,000**      c) 102.45 **102**      d)  $2.301 \times 10^{-2}$   **$2.30 \times 10^{-2}$**
- Calculate, use sig figs:  $(12.22 \times 14.02 \times 120.1 \times .555) / 10.00 =$  **1140**
- A sample of metal having a mass of 12.122 g is added to a graduated cylinder containing 35.2 mL of water. The water level rises to 38.4mL. Calculate the density, use correct significant figures.  
$$d = \frac{12.122}{3.2} = 3.8 \text{ g/mL}$$
- Calculate the % error for an experiment where the student measured 3.45 g and the actual value was 3.00 g?  
$$\% \text{ error} = \frac{(3.45 \text{ g} - 3.00 \text{ g})}{(3.00 \text{ g})} \times 100 = 15.0\%$$
- Write the symbol for the following ions: Mg losing 2 e<sup>-</sup>, Al losing 3 e<sup>-</sup>, N gaining 3 e<sup>-</sup>, S gaining 2 e<sup>-</sup>  
**Mg<sup>2+</sup>, Al<sup>3+</sup>, N<sup>3-</sup>, S<sup>2-</sup>**
- a) 2 b) 3 c) 12 d) 5
- List three points of Dalton's atomic theory.  
A. Elements made of tiny particles, called atoms.  
B. Atoms of a given element are identical; the atoms of different elements are different.  
C. In chemical reactions, atoms are combined, separated, or rearranged. Atoms cannot be subdivided, created, or destroyed.  
D. Atoms combine with atoms, in small whole numbers, to form compounds.
- Describe Thompson's experiment with the cathode ray tube, what did he discover?  
Cathode ray tube consists of a cathode (negative) and anode (positive) side. The "ray" could be bent with magnetic or electric fields. The "ray" bent toward a positive EM field and away from a negative EM field. Discovered electron and charge to mass ratio of electron.
- Describe Millikan's Oil Drop experiment. What exactly did he determine?  
He suspended a micro oil droplet in an electric field and was able to determine the actual charge of an electron to be  $-1.6 \times 10^{-19} \text{ C}$ . Also, determined the mass of an electron.
- Explain Rutherford's Gold Foil experiment. What were three observations and three conclusions?  
He fired alpha particles at gold foil and observed: 1) most went straight through, 2) small number showed a slight deflection, 3) very few bounced straight back. He concluded: 1) mass of atom is in the nucleus, 2) nucleus is positively charged, 3) atom mostly empty space.
- What is an isotope? **Atoms of the same element that differ by # of electrons.**
- Write the symbol of the most abundant copper isotope if the average atomic mass is 63.55?  **$^{64}_{29}\text{Cu}$**
- Write the nuclear symbol for an isotope with Z=40, and 45 neutrons?  **$^{85}_{40}\text{Zr}$**
- Consider the diagram and match the letters with the appropriate particle  
**A= electron, B= neutron, C= proton**
- Place the forms of hydrogen (H, H<sup>+</sup>, H<sup>-</sup>, H<sub>2</sub>) in the table.



Atom	Cation	Anion	Molecule
<b>H</b>	<b>H<sup>+</sup></b>	<b>H<sup>-</sup></b>	<b>H<sub>2</sub></b>

# Modeling Chemistry Unit 1 Review Answers

**Guillaume Favre**



## Modeling Chemistry Unit 1 Review Answers:

*Jacaranda Chemistry 1 VCE Units 1 and 2, learnON and Print* Neale Taylor, Angela Stubbs, Robert Stokes, 2022-11-30  
Developed by expert Victorian teachers for VCE students The NEW Jacaranda Chemistry VCE series continues to deliver curriculum aligned material that caters to students of all abilities Our expert author team of practising teachers and assessors ensures 100% coverage of the new VCE Chemistry Study Design 2023 2027      **Sif Chemistry NI Tb** Rex M. Heyworth, 2007      **Chemistry in the Community.** American Chemical Society, 2002 This volume has relevance to a wide number of courses giving a hands on introduction to chemistry in relation to community issues rather than around specific chemical concepts      **Shreir's Corrosion** , 2009-02-27 This four volume reference work builds upon the success of past editions of Elsevier's Corrosion title by Shreir Jarman and Burstein covering the range of innovations and applications that have emerged in the years since its publication Developed in partnership with experts from the Corrosion and Protection Centre at the University of Manchester Shreir's Corrosion meets the research and productivity needs of engineers consultants and researchers alike Incorporates coverage of all aspects of the corrosion phenomenon from the science behind corrosion of metallic and non metallic materials in liquids and gases to the management of corrosion in specific industries and applications Features cutting edge topics such as medical applications metal matrix composites and corrosion modeling Covers the benefits and limitations of techniques from scanning probes to electrochemical noise and impedance spectroscopy

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