

## Unit 6 – Worksheet 3

### Ionic Compounds

#### Properties

Ionic compounds generally have higher m.p's and b.p's than do molecular compounds. When molten, they conduct electricity. If soluble, the solutions also conduct electricity.

#### Basic structural unit

The basic structural unit is the formula unit, the simplest ratio of atoms in a crystalline ionic compound.

1. Give the name of the following simple binary ionic compounds.

- a.  $\text{Na}_2\text{O}$  sodium oxide
- b.  $\text{K}_2\text{S}$  potassium sulfide
- c.  $\text{MgCl}_2$  magnesium chloride
- d.  $\text{CaBr}_2$  calcium bromide
- e.  $\text{BaI}_2$  barium iodide
- f.  $\text{Al}_2\text{S}_3$  aluminum sulfide
- g.  $\text{CaBr}$  cesium bromide
- h.  $\text{AgF}$  silver fluoride

2. Give the name of the following simple binary ionic compounds.

- a.  $\text{Na}_3\text{N}$  sodium nitride
- b.  $\text{K}_2\text{O}$  potassium oxide
- c.  $\text{AgBr}$  silver bromide
- d.  $\text{MgI}_2$  magnesium iodide
- e.  $\text{SrO}$  strontium oxide

3. Write the formula for the following binary ionic compounds.

- a. lithium bromide  $\text{LiBr}$
- b. sodium iodide  $\text{NaI}$
- c. silver sulfide  $\text{Ag}_2\text{S}$
- d. cesium oxide  $\text{Cs}_2\text{O}$
- e. beryllium iodide  $\text{BeI}_2$
- f. barium hydride  $\text{BaH}_2$
- g. aluminum fluoride  $\text{AlF}_3$
- h. potassium oxide  $\text{K}_2\text{O}$

4. Write the formula for these ionic substances.

- a. silver oxide  $\text{Ag}_2\text{O}$
- b. aluminum sulfide  $\text{Al}_2\text{S}_3$
- c. sodium nitride  $\text{Na}_3\text{N}$
- d. barium chloride  $\text{BaCl}_2$
- e. strontium hydride  $\text{SrH}_2$

# Modeling Chemistry U6 Wsv2 0 Answers

**Raffaela Di Napoli**



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