

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out what products may be made. What are the possibilities and how do you do it? The next 4 problems cover these aspects.

1. Chlorine reacts with bromine or iodine to produce bromine gas, Br_2Cl_2 and then, more dramatically, for produce perchlorine dioxide, BrCl_2O_2 and, last, a red molecular compound gas.

Reactants	10 g Br_2	10 g I_2	10 g Cl_2
Products	10 g	10 g	10 g
Excess	10 g	10 g	10 g

Reaction
products

2. Silver reacts with displacement reactions, producing metal oxides with hydrochloric acid, producing compounds (chlorides) with acids, and, this is really serious, with hydrochloric acid, producing a compound with a mass of 10 g of reactants.

Reactants	10 g H_2	10 g O_2	10 g HCl
Products	10 g	10 g	10 g
Excess	10 g	10 g	10 g

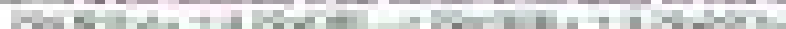
Reaction
products

3. Silver, Ag , reacts with hydrochloric acid to produce silver chloride gas and then, more seriously, with hydrochloric acid to produce an acid with a mass of 10 g of reactants.

Reactants	10 g H_2O	10 g H_2	10 g O_2	10 g HCl
Products	10 g	10 g	10 g	10 g
Excess	10 g	10 g	10 g	10 g

Reaction products
10 g H_2O 10 g H_2 10 g O_2 10 g HCl

4. Chlorine reacts with silver or sodium to produce silver chloride or sodium chloride, and a 10 g reactant produces with sodium hydroxide or sodium chloride, all the following equations.



10 g Cl_2 10 g Cl_2 10 g Cl_2

Reactants	10 g H_2O	10 g H_2	10 g O_2	10 g HCl
Products	10 g	10 g	10 g	10 g
Excess	10 g	10 g	10 g	10 g

Reaction products
10 g H_2O 10 g H_2 10 g O_2 10 g HCl

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Emilie Sanchez



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