

Chemistry – Unit 3

Energy and Kinetic Molecular Theory

In the 18th and 19th centuries scientists wrestled with identifying and describing the nature of the “stuff” that produced change. One concept that became popular for a while was that of “caloric” (what we now call heat).

“Caloric was originally conceived of as a quantity that would flow from a hotter object to a cooler one that would warm up as a result. It answered the need for a way for the cause of warming to get from here to there. Not only did caloric serve as a cause for warming, it was also considered to be the cause for changes of phase. Caloric enabled particles of a substance to move farther apart until the attraction of the particles for each other became too weak to hold them together. Although Lavoisier did not think that caloric necessarily was an actual substance, in its storage and transfer it was *like* a substance.”¹

When scientists recognized that the “stuff” involved when forces were applied to objects to lift them or change their speed was the same “stuff” that was involved when the temperature of objects changed, they worked to develop a single energy concept. “So when the energy concept was developed it was important to distinguish it from caloric. In snuffing out the caloric concept, the clear picture of energy storage and transfer that it fostered was unnecessarily lost, too.”²

Even though we recognize that energy is not a *physical* substance, we choose to use the substance *metaphor* to describe it.

We'll use three principles to guide us in the development of the energy concept.

1. Energy can be viewed as a substance-like quantity that can be stored in a physical system.
2. Energy can “flow” or be “transferred” from one system to another and so cause changes.
3. Energy maintains its identity after being transferred.

If you are unsure what we mean by the use of a substance metaphor, consider how we describe information. We say that it can be stored in books, on computer hard drives or floppy disks or CD-ROMs. Information can be transferred from place to place via cables or by wireless transmission techniques - in fact you just did this when you accessed this lesson via the Internet, transferred it to your computer and then (perhaps) printed it. But there is nothing substantial about the information itself; you can't touch it or measure its mass on a balance. The third point is important to consider because many texts talk about energy *transformations* as if somehow it is the *energy* that is changing rather than the physical system that gains or loses it. Consider the information metaphor again; even though we move information from place to place or store it in different ways, nothing about the information itself has changed.

Modeling Chemistry U3 Reading V2 0 Answers

Michael Seilmaier



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Table of Contents Modeling Chemistry U3 Reading V2 0 Answers

1. Understanding the eBook Modeling Chemistry U3 Reading V2 0 Answers
 - The Rise of Digital Reading Modeling Chemistry U3 Reading V2 0 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U3 Reading V2 0 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry U3 Reading V2 0 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U3 Reading V2 0 Answers
 - Personalized Recommendations

- Modeling Chemistry U3 Reading V2 0 Answers User Reviews and Ratings
- Modeling Chemistry U3 Reading V2 0 Answers and Bestseller Lists
- 5. Accessing Modeling Chemistry U3 Reading V2 0 Answers Free and Paid eBooks
 - Modeling Chemistry U3 Reading V2 0 Answers Public Domain eBooks
 - Modeling Chemistry U3 Reading V2 0 Answers eBook Subscription Services
 - Modeling Chemistry U3 Reading V2 0 Answers Budget-Friendly Options
- 6. Navigating Modeling Chemistry U3 Reading V2 0 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry U3 Reading V2 0 Answers Compatibility with Devices
 - Modeling Chemistry U3 Reading V2 0 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry U3 Reading V2 0 Answers
 - Highlighting and Note-Taking Modeling Chemistry U3 Reading V2 0 Answers
 - Interactive Elements Modeling Chemistry U3 Reading V2 0 Answers
- 8. Staying Engaged with Modeling Chemistry U3 Reading V2 0 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry U3 Reading V2 0 Answers
- 9. Balancing eBooks and Physical Books Modeling Chemistry U3 Reading V2 0 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry U3 Reading V2 0 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry U3 Reading V2 0 Answers
 - Setting Reading Goals Modeling Chemistry U3 Reading V2 0 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry U3 Reading V2 0 Answers
 - Fact-Checking eBook Content of Modeling Chemistry U3 Reading V2 0 Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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