

Solubility Rules for Aqueous Solutions

“Sol.” means that more than 3g of the substance dissolves in 100 ml of water. “Ppt.” indicates that the combination forms a precipitate.

| Solubility of Aqueous Solutions | | | | | | | |
|--|----------------------------|------------------|--------|------|------|---------------|-----------------|
| | alkali or NH_4 | Ag, Hg, or Pb | Ba, Sr | Ca | Mg | Fe, Cu, Zn | other metals |
| nitrate (NO_3^-), acetate (CH_3COO^-), chlorate (ClO_3^-), perchlorate (ClO_4^-) | sol. | sol. | sol. | sol. | sol. | sol. | sol. |
| fluoride (F^-) | sol. | ppt. | ppt. | ppt. | ppt. | ppt. | ppt. |
| chloride (Cl^-), bromide (Br^-), iodide (I^-) | sol. | ppt. | sol. | sol. | sol. | sol. | sol. |
| sulfate (SO_4^{2-}) | sol. | ppt. | ppt. | ppt. | sol. | sol. | sol. |
| carbonate (CO_3^{2-}), phosphate (PO_4^{3-}) | sol. | ppt. | ppt. | ppt. | ppt. | ppt. | ppt. |
| oxide (O^{2-}) | form OH | ppt. | ppt. | ppt. | ppt. | ppt. | ppt. |
| hydroxide (OH^-), sulfide (S^{2-}) | sol. | ppt. | sol. | sol. | ppt. | ppt. | ppt. |
| chromate (CrO_4^{2-}) | sol. | ppt. | ppt. | sol. | sol. | ppt. | ppt. |

Compounds that are Soluble in Water

1. All common salts of the alkali metals (Group IA/1) or ammonium (NH_4^+) are soluble.
2. All common acetates (CH_3COO^-) and nitrates (NO_3^-) are soluble.
3. All binary compounds of a metal plus a halogen (Group VIIA/17) are soluble except for fluorides (F^-) and compounds containing silver (Ag), mercury (Hg), or lead (Pb).
4. All sulfates (SO_4^{2-}) are soluble except for those of barium (Ba), strontium (Sr), lead (Pb), calcium (Ca), silver (Ag), or mercury (Hg).

Compounds that are Insoluble in Water

1. All carbonates (CO_3^{2-}), phosphates (PO_4^{3-}), and oxides are insoluble except for alkalis (soluble compounds rule #1).
2. All hydroxides (OH^-) and sulfides are insoluble except for alkalis (soluble compounds rule #1) and compounds containing calcium (Ca), strontium (Sr), or barium (Ba).
3. All chromates (CrO_4^{2-}) are insoluble except for alkalis (soluble compounds rule #1) and compounds containing calcium (Ca) or magnesium (Mg).

Model Rules Of Solubility In Aqueous Solutions Answers

Bahl Arun/ Bahl B.S. & Tuli G.D.



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Supplement to Mellor's Comprehensive Treatise on Inorganic and Theoretical Chemistry Joseph William Mellor, 1922 **An Introduction to Chemistry** Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

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