

COURSE STRUCTURE**Class XI****One Paper****Three Hours****Max Marks. 100**

Units	Marks
I. SETS AND FUNCTIONS	29
II. ALGEBRA	37
III. COORDINATE GEOMETRY	13
IV. CALCULUS	06
V. MATHEMATICAL REASONING	03
VI. STATISTICS AND PROBABILITY	12

UNIT-I: SETS AND FUNCTIONS**1. Sets :****(12) Periods**

Sets and their representations. Empty set. Finite & Infinite sets. Equal sets. Subsets. Subsets of a set of real numbers especially intervals (with notations). Power set. Universal set. Venn diagrams. Union and Intersection of sets. Difference of sets. Complement of a set. Properties of Complement Sets.

2. Relations & Functions:**(14) Periods**

Ordered pairs, Cartesian product of sets. Number of elements in the cartesian product of two finite sets. Cartesian product of the set of reals with itself (upto $\mathbb{R} \times \mathbb{R} \times \mathbb{R}$). Definition of relation, pictorial diagrams, domain, codomain and range of a relation. Function as a special kind of relation from one set to another. Pictorial representation of a function, domain, co-domain & range of a function. Real valued functions, domain and range of these functions, constant, identity, polynomial, rational, modulus, signum and greatest integer functions, with their graphs. Sum, difference, product and quotients of functions.

3. Trigonometric Functions:**(18) Periods**

Positive and negative angles. Measuring angles in radians & in degrees and conversion from one measure to another. Definition of trigonometric functions with the help of unit circle. Truth of the identity $\sin^2 x + \cos^2 x = 1$, for all x . Signs of trigonometric functions. Domain and range of trigonometric functions and their graphs. Expressing $\sin(x \pm y)$ and $\cos(x \pm y)$ in terms of $\sin x$, $\sin y$, $\cos x$ & $\cos y$. Deducing the identities like the following:

$$\tan(x \pm y) = \frac{\tan x \pm \tan y}{1 \mp \tan x \tan y}, \cot(x \pm y) = \frac{\cot x \cot y \mp 1}{\cot y \pm \cot x},$$

$$\sin x + \sin y = 2 \sin \frac{x+y}{2} \cos \frac{x-y}{2}, \cos x + \cos y = 2 \cos \frac{x+y}{2} \cos \frac{x-y}{2},$$

$$\sin x - \sin y = 2 \cos \frac{x+y}{2} \sin \frac{x-y}{2}, \cos x - \cos y = -2 \sin \frac{x+y}{2} \sin \frac{x-y}{2},$$

Identities related to $\sin 2x$, $\cos 2x$, $\tan 2x$, $\sin 3x$, $\cos 3x$ and $\tan 3x$. General solution of trigonometric equations of the type $\sin \theta = \sin \alpha$, $\cos \theta = \cos \alpha$ and $\tan \theta = \tan \alpha$. Proof and simple applications of sine and cosine formulae.

UNIT-II: ALGEBRA**1. Principle of Mathematical Induction:****(06) Periods**

Process of the proof by induction, motivating the application of the method by looking at natural numbers as the least inductive subset of real numbers. The principle of mathematical induction and simple applications.

Maths P2014 Grade11

**Ann Kajander, Jennifer Holm, Egan J
Chernoff**



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web 1 78 348 stoichiometry helps us determine how much substance is needed or is present things that can be measured are reactants and products mass molecular weight chemical equations formulas stoichiometric coefficient the stoichiometric coefficient or stoichiometric number is the number of molecules that participate in the reaction

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$$\begin{array}{l} 1 \text{ N}_2 + \text{F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} \\ 3 \text{ HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \\ \text{GaBr} + \text{NaBr} + \text{SnO}_4 \rightarrow \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + 3 \text{ NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3 \end{array}$$
 using the equation from problem 2 above answer the following

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$$\begin{array}{l} 1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O} \\ 3 \text{ HBr} + 1 \text{ KHCO}_3 \rightarrow 1 \text{ H}_2\text{O} + 1 \text{ KBr} + \text{CO}_2 \\ 2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 \rightarrow 1 \text{ Ga}_2\text{SO}_3 + 3 \text{ NaBr} + 3 \text{ SnO}_2 + 3 \text{ NF}_3 + \text{SnF}_2 + 1 \text{ N}_2\text{O}_3 \end{array}$$
 using the following equation $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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assuming ce cucl_2 cucx_2 is available in excess

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 $\text{br}_1 + \text{n}_2 \rightarrow \text{f}_2 + \text{nf}_3$
 $\text{br}_2 + \text{c}_6\text{h}_{10} + \text{o}_2 \rightarrow \text{co}_2 + \text{h}_2\text{o}$
 $\text{br}_3 + \text{hbr} + \text{khco}_3 \rightarrow \text{h}_2\text{o} + \text{kbr} + \text{co}_2$
 $\text{br}_4 + \text{gabr}_3 + \text{na}_2\text{so}_3 \rightarrow \text{ga}_2\text{so}_3 + \text{nabr}$
 $\text{br}_5 + \text{sno}_2 + \text{nf}_3 \rightarrow \text{snf}_2 + \text{n}_2\text{o}_3$
 using the following equation
 $\text{br}_2 + \text{naoh} + \text{h}_2\text{so}_4 \rightarrow \text{h}_2\text{o} + \text{na}_2\text{so}_4$

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 $\text{c}_6\text{h}_{10} + \text{o}_2 \rightarrow \text{co}_2 + \text{h}_2\text{o}$
 $\text{hbr} + \text{khco}_3 \rightarrow \text{h}_2\text{o} + \text{kbr} + \text{co}_2$
 $\text{gabr} + \text{na}_2\text{so}_3 \rightarrow \text{ga}_2\text{so}_3 + \text{nabr}$
 $\text{sno}_2 + \text{nf}_3 \rightarrow \text{snf}_2 + \text{n}_2\text{o}_3$

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 $\text{ce sio}_2 \rightarrow \text{sic}$

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 $\text{n}_2 + \text{f}_2 \rightarrow \text{nf}_3$
 $\text{c}_6\text{h}_{10} + \text{o}_2 \rightarrow \text{co}_2 + \text{h}_2\text{o}$
 $\text{hbr} + \text{khco}_3 \rightarrow \text{h}_2\text{o} + \text{kbr} + \text{co}_2$
 $\text{gabr} + \text{na}_2\text{so}_3 \rightarrow \text{ga}_2\text{so}_3 + \text{nabr}$
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same thing goes with the o3 i am so lost

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web google classroom balance the following chemical equation $\text{Mg} + \text{OH}^- \rightarrow \text{Mg}(\text{OH})_2$ $\text{HCl} \rightarrow \text{HCl}$ MgCl_2 H_2 O H_2O note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem $7.41 \times 10^5 \text{ y}$ θ π 8.52×10^6

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