

Isotopes Practice Set

- What does atomic number tell you about an element?
- What does the number used to isotopes signify?
- How can you tell isotopes apart?

For each of the following isotopes, write the number of protons, neutrons, and electrons.

	Chlorine-35	Chlorine-37
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Nitrogen-14	Nitrogen-15
# of protons	7	7
# of neutrons	7	8
# of electrons	7	7

	Carbon-12	Carbon-13
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Oxygen-16	Oxygen-18
# of protons	8	8
# of neutrons	8	10
# of electrons	8	8

Fill in the isotopes names including chemical symbol, atomic number, and mass number from the values that your partner gave and the information provided.

	Isotope: $^{35}_{17}\text{Cl}$	Isotope: $^{37}_{17}\text{Cl}$
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Isotope: $^{12}_6\text{C}$	Isotope: $^{13}_6\text{C}$
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Isotope: $^{14}_7\text{N}$	Isotope: $^{15}_7\text{N}$
# of protons	7	7
# of neutrons	7	8
# of electrons	7	7

	Isotope: $^{16}_8\text{O}$	Isotope: $^{18}_8\text{O}$
# of protons	8	8
# of neutrons	8	10
# of electrons	8	8

Isotope Practice Set Answers

R Sanford

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