



Ink Bridge Study Guide

D Kirk



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parchment scripts decoration illustration and bindings mix with consideration of such conceptual categories as ownership authorship language miscellaneity geography writing editing mediation illustration and performance as well as of the status of the literary itself Each case study includes an essay orienting the reader to particularly productive categories of analysis and a selected bibliography for further research Because a high quality digital surrogate exists for each of the selected manuscripts fully and freely available online readers can gain access to the artifacts in their entirety enabling further individual exploration and facilitating the book's classroom use Medieval English Manuscripts and Literary Forms aims to inspire a broad group of readers with some of the excitement of literary manuscript studies in the twenty first century The interpretative frameworks surrounding each object will assist everyone in thinking through the implications of manuscript culture more generally not only for the deeper study of the literature of the Middle Ages but also for a better understanding of book cultures of any era including our own

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web donor whereas a base generically rep resented by b is a proton acceptor when acids and bases combine they swap a proton ha b a hb strong acids e g sulfuric acid have a very pronounced ten dency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid $H_2C_2O_4 \cdot 2H_2O$ was weighted to the nearest 0.01 g in a 50 mL beaker 2 about

30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0 100 m the ph of the solution is then $\text{ph} = -\log 0 100 = 1 000$ $\text{ph} = -\log 0 100 = 1 000$ b titrant volume 12 50 ml

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to determine the concentration of an acid with a base to determine the concentration of an acidic sample

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