

- Carbohydrates contain carbon, hydrogen and oxygen, usually with the general formula $C_nH_{2n}O_n$

Glucose

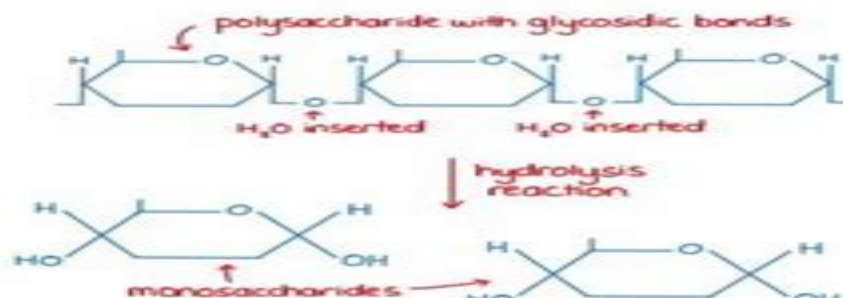
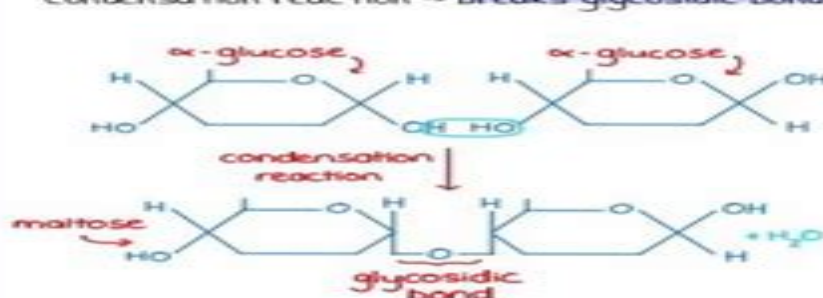
- A monosaccharide with the formula $C_6H_{12}O_6$
- A hexose monosaccharide (six carbon atoms) in a ring structure
- Soluble in water → easily transported
- Main energy source for animals and plants
→ chemical bonds store lots of energy
- Two isomers: α -glucose and β -glucose → H and OH groups on carbon 1 inverted in β -glucose

Monosaccharides are small soluble carbohydrate monomers. They also include fructose and galactose.



Glycosidic bonds and condensation/hydrolysis reactions

- Condensation reaction: two molecules join to form a new chemical bond and a water molecule is eliminated
- Condensation reactions form glycosidic bonds between monosaccharides to create disaccharides and polysaccharides
- Hydrolysis reaction: a water molecule is used and the chemical bond is broken (reverse of a condensation reaction → breaks glycosidic bonds)



Disaccharides

- Two monosaccharides joined together with a glycosidic bond in a condensation reaction
- Soluble in water
- Maltose = glucose + glucose
- Sucrose = glucose + fructose
- Lactose = glucose + galactose

Monosaccharides and disaccharides are sugars.

Polysaccharides

- Large polymers of monosaccharides joined with glycosidic bonds
- Starch and glycogen are large energy storage molecules which cannot leave cells

Starch

- Glucose storage in plants → hydrolysed when glucose is needed
- Insoluble in water → does not affect the water potential of cells so water is not drawn in by osmosis
- Amylose → unbranched α -glucose polysaccharide (1,4-glycosidic bonds)

Iodine test for starch

- Add iodine in potassium iodide solution to sample.
- If starch is present: goes from brown/orange to blue-black.

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