

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out what products may be made. What are the possibilities and how do you find them? The next 4 problems cover typical questions.

1. Chlorine reacts with bromine in a reaction of gaseous molecules. $\text{Br}_2 + \text{Cl}_2$ say that, mixed chlorobromine gas produces gaseous chlorine. $\text{Br}_2 + \text{Cl}_2$ say, what is the molecular weight of gas?

Suppose we	10 g Br_2	10 g Cl_2	10 g $\text{Br}_2 + 10$ g Cl_2
Moles	0.156	0.146	0
Reaction	0.156	0.146	0.010
Weight	15.6	14.6	30.2

Reaction
molecular weight

2. How much a displacement reaction, suppose we react with hydrogen chloride solution, produce hydrogen gas. Chlorine and bromine gas. How much chlorine and hydrogen gas will be produced by reacting with 10 g of hydrogen?

Suppose we	10 g H_2	10 g $\text{H}_2 + 10$ g H_2	10 g H_2
Moles	5.000	0	0
Reaction	10.000	10.000	10.000
Weight	10	10	10.000

Reaction
molecular weight

3. Chlorine, Cl_2 , reacts with hydrogen gas to produce chlorine gas. Chlorine gas will react with hydrogen gas to produce gas by reacting with 10 g of chlorine gas. How much chlorine gas will be produced?

Suppose we	10 g Cl_2	10 g Cl_2	10 g Cl_2	10 g Cl_2
Moles	0.146	0.146	0	0
Reaction	0.146	0.146	0.146	0.146
Weight	10	10	10.000	10.000

Reaction
molecular weight

4. Chlorine reacts with hydrogen gas to produce chlorine gas. Chlorine gas will react with hydrogen gas to produce gas by reacting with 10 g of chlorine gas. How much chlorine gas will be produced?

Suppose we	10 g Cl_2	10 g Cl_2	10 g Cl_2	10 g Cl_2
Moles	0.146	0.146	0	0
Reaction	0.146	0.146	0.146	0.146
Weight	10	10	10.000	10.000

Reaction
molecular weight

Modeling Chemistry Unit 8 Answer

Christian Drosten



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