

Chemistry Unit 7 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedures outlined in class. Be sure to circle your final answer.

- Calculate the number of moles of potassium chlorate, KClO_3 , that must decompose to produce potassium chloride, KCl , and 1.0 moles of oxygen gas.
- In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 0.40 g of magnesium?
- Ethanol, $\text{C}_2\text{H}_5\text{OH}$, reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.00 moles of ethanol?
- Determine the mass of sodium chloride produced when 0.70 g of sodium (Na) reacts with sodium hydroxide according to the following equation:
 $\text{NaClO}_3 + 2 \text{NaOH} \rightarrow \text{NaClO}_2 + 2 \text{NaOH}$
- In the copper-silver nitrate lab copper metal and silver nitrate solution reacted to produce silver metal and copper(II) nitrate in solution.
 A student placed a copper wire with a mass of 0.05 g in the reaction test tube. The silver-silver nitrate solution contained 0.40 g of silver nitrate.
 She obtained 0.07 g of silver metal. Calculate the percent yield of silver.
- When hydrochloric acid (HCl) is added to sodium hydroxide solution, the products are water, aqueous sodium chloride and sodium chloride gas. What is the per cent yield of 0.04 g of Cl_2 gas collected when 0.01 g of sodium hydroxide solution reacts with excess HCl?
- Phosphorus and bromine react (chemically) together to form phosphorus tribromide. If 0.01 g of phosphorus and 0.01 g of bromine react, how many grams of PBr_3 could be produced?
- Iron sulfate and oxygen gas react to form iron oxide and sulfur dioxide. Determine the amount of SO_2 that should be produced in a reaction between 0.01 g of FeSO_4 and 0.01 g of oxygen. What is the percent of the experiment?

$$1) 1.2 \text{ moles of } \text{KClO}_3 \quad 2) 1.000 \text{ moles of } \text{HCl} \quad 3) 0.001 \text{ g } \text{NaCl} \quad 4) 0.000 \text{ g of } \text{NaClO}_2$$

$$5) 0.01\% \quad 6) 100.0\% \quad 7) 0.01 \text{ g } \text{PBr}_3 \quad 8) 0.001 \text{ g of } \text{SO}_2 \quad 9) 0.01 \text{ g of } \text{FeSO}_4$$

Modeling Chemistry Unit 7 Ws 4 V2

Honghua Tan



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Comprehensive Biomaterials II Kevin Healy, Dietmar W. Hutmacher, David W. Grainger, C. James

Kirkpatrick, 2017-05-18 Comprehensive Biomaterials II Second Edition Seven Volume Set brings together the myriad facets of biomaterials into one expertly written series of edited volumes. Articles address the current status of nearly all biomaterials in the field, their strengths and weaknesses, their future prospects, appropriate analytical methods and testing device applications and performance, emerging candidate materials as competitors and disruptive technologies, research and development, regulatory management, commercial aspects and applications including medical applications. Detailed coverage is given to both new and emerging areas and the latest research in more traditional areas of the field. Particular attention is given to those areas in which major recent developments have taken place. This new edition with 75% new or updated articles will provide biomedical scientists in industry, government, academia and research organizations with an accurate perspective on the field in a manner that is both accessible and thorough. Reviews the current status of nearly all biomaterials in the field by analyzing their strengths and weaknesses, performance and future prospects. Covers all significant emerging technologies in areas such as 3D printing of tissues, organs and scaffolds, cell encapsulation, multimodal delivery, cancer vaccine, biomaterial applications, neural interface, understanding materials used for in situ imaging and infection prevention and treatment. Effectively describes the many modern aspects of biomaterials from basic science to clinical applications.

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are now available in a hardcover print format as well as a special electronic edition This volume provides valuable content for all researchers in theoretical chemistry and will especially benefit those research groups and libraries with limited access to the journal

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