

Answer all the questions in the spaces provided.

- 1 (a) The solubility of the Group 2 sulfates decreases down the group.

Explain this trend.

.....

.....

.....

.....

..... [3]

- (b) Describe what is observed when magnesium and barium are reacted separately with an excess of dilute sulfuric acid.

magnesium

barium

[1]

- (c) The solubility product, K_{sp} , of BaSO_4 is $1.08 \times 10^{-10} \text{ mol}^2 \text{ dm}^{-6}$ at 298 K.

Calculate the solubility of BaSO_4 in g per 100 cm^3 of solution.

solubility of BaSO_4 = g per 100 cm^3 of solution [2]

Modeling Chemistry U6 Test

R Barnett



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