

**Inquiry Question**

**How do we determine the solubility product equilibrium constant ( $K_{sp}$ ) for lead (II) iodide ( $PbI_2$ )?**

Name: Diego MeznickDate: 2022-07-05

Up until now, we have looked at reactions that can achieve equilibrium. In Unit 3 (Solubility), you look at a special case of equilibrium where one of the products formed may form a solid. This process is an equilibrium process.

Now, for some clarity, let's look at an example. Let's show a balanced chemical reaction of lead (II) nitrate with potassium iodide.



You will notice a solid is one of the products. Even though the reaction is lead (II) nitrate with potassium iodide, we are shifting our "spotlight" to look only at the lead (II) iodide product. It is this product and the equilibrium it forms between its solid form and its dissociated ion form.

Understanding the degree to which the solid will dissolve and dissociate compared to other products can help us predict outcomes in experiments or even help drive outcomes in specific directions (ex. choosing reactants whose products display very low solubility).

During this activity, you will determine the solubility product constant ( $K_{sp}$ ) for  $PbI_2$  by watching a video of an experiment.

Begin by watching this video: [Determination of a Solubility Product Constant](#)

# Modeling Chemistry U4 Lab V2 0

**Sasha Hu**



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