At a high temperature, 0.50 mol of HBr was placed in a 1.0 L container and allowed to decompose according to the reaction:

$$2HBr(g) \rightleftharpoons H_{2(g)} + Br_{2(g)}$$

At equilibrium the [Br₂] was measured to be 0.13 M. What is K_{eq} for this reaction at this temperature?

	2HBr =	= H2 +	Br2
[x]	0.50	0	0
[C]	-0-26	+0.13	+0.13
IE]	0.24	0.13	0.13

Keg =	IH2][Br]2	= (0.13)2	=0.29
	100001100000000000000000000000000000000	150000000000000	

Answer Keg = 0.29

 When 1.0 mol of NH_{3(g)} and 0.40 mol of N_{2(g)} are placed in a 5.0 L vessel and allowed to reach equilibrium at a certain temperature, it is found that 0.78 mol of NH₃ is present. The reaction is:

2NH3(g) = 3H2(g) + N2(g)

Initial [NH3] = 1.0 mol = 0.20 M Initial [N2] = 0.40 mol = 0.080 M

Equil [NH3] = 0.78 mol = 0.156 M 2NH2 = 3 H2 + N2

	aNH3 &	= 3 Na	+ N2
[I]	0.20	0	0.080
EC3	-0.044	+0.066	+0.022
LEJ.	0.156	0.066	0.102

a) Calculate the equilibrium concentrations of all three species.

[NH₃] - 0.16 M [H₂] - 0.066 M [N₂] - 0.10 M

b), Calculate the value of the equilibrium constant at this temperature. (use unrounded cones, then round to 2503)

 $\text{Keq} = \frac{\text{EH}_2 \text{J}^3 \text{EN}_2 \text{J}}{\text{ENH}_3 \text{J}^2} = \frac{(0.066)^3 (0.102)}{(0.156)^2} = 0.0012$ Answ

Answer 1.2×10-3

c) How many moles of H₂ are present at equilibrium?

0.066 M x 5.0 L = 0.33 mal

Answer 0.33 mol

d) How many moles of N₂ are present at equilibrium?

0.102 M x 5.0 L = 0.51 mal

Answer 0.51 mol

Modeling Chemistry Ws Answers

Christian Drosten

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In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has be more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Modeling Chemistry Ws Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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