

$$\begin{aligned}
 K' &= \frac{[\text{H}_2\text{O}_2]}{[\text{NO}_2]^2} \\
 &= \frac{(9.3 \times 10^{-3})}{[\text{NO}_2]^2} \\
 &= \frac{(9.3 \times 10^{-3})}{(66.5 \times 10^{-3} \text{ mol/L})^2} \\
 &= 2.20 \\
 &= 2.2 \times 10^2
 \end{aligned}$$

45. (a) Students should expect it to be a base.



- (c) Candy contains sucrose, which will remove hydroxyapatite from teeth by reacting with it at low pH to give  $\text{Ca}^{2+}_{(aq)}$  and  $\text{H}_2\text{PO}_4^-_{(aq)}$  that are in soluble form.

46.

	$\text{NH}_3 \cdot (\text{CH}_3)_2\text{NH}_2$	$\text{NH}_3^+ \cdot (\text{CH}_3)_2\text{NH}_2^+$	$\text{OH}^-$
initial conc.	0.10	0	0
change	-0.0021	+0.0021	+0.0021
equilibrium	0.0979	0.0021	0.0021

$$K_b = \frac{(0.0021)^2}{0.0979} = 4.5 \times 10^{-5}$$

47. The conductivity of 0.1 mol/L HCl is higher than for 0.1 mol/L  $\text{CH}_3\text{COOH}$  because the HCl is completely ionized whereas the  $\text{CH}_3\text{COOH}$  is only partly ionized. For solutions of these two acids at a concentration of  $1 \times 10^{-7}$  mol/L, the HCl is still ionized completely and the  $\text{CH}_3\text{COOH}$  again is only partly ionized. However at such a low concentration the conductivity of the HCl is the same as for water. There will be no discernable difference in the conductivities at this low concentration.

48. One drop of concentrated acid will have a high concentration of  $\text{H}_3\text{O}^+$  ions. (E.g., For HCl  $[\text{H}_3\text{O}^+] =$  about 12.5 mol/L, for  $\text{H}_2\text{SO}_4$   $[\text{H}_3\text{O}^+] =$  about 35 mol/L. When one drop of acid is added to a bucket of water, the  $\text{H}_3\text{O}^+$  ion concentration is diluted by a very large factor.

49. (a) A triprotic acid is an acid that has three protons that may be dissociated, each with its own  $K_a$  value.

- (b) Stepwise Equation 1:



- Stepwise Equation 2:



- Stepwise Equation 3:



- (c)  $\text{H}_2\text{PO}_4^-_{(aq)}$  acts as an acid in equation 2 above. In the equation below, it acts as a base.



40. Students may say that  $\text{H}_3\text{PO}_{4(laq)}$  is the stronger acid because it is higher in the list of relative strengths of acids and bases. Some students may suggest that  $\text{H}_3\text{PO}_{4(laq)}$  is a stronger acid because its  $K_a$  value is larger than that of  $\text{H}_2\text{PO}_4^-_{(aq)}$ , which is a "stronger" answer.

50. Water could have a pH greater than 7.0 at a temperature different from 25°C.

51. (a) Students might suggest either of these equations:



- (b) The litmus test indicates that the solution is basic, so the solution contains a considerable  $[\text{OH}^-]$ . Neither of these equations, written according to the original Arrhenius theory, is consistent with the litmus test. Equation two above even implies a red litmus result! Accordingly, the litmus test contradicts the Arrhenius theory equations.

- (c) A student could have two frames of thought.

Perspective I:



or, without the spectator,  $\text{Na}^+$ :



Perspective II



or, without the spectator,  $\text{Na}^+$ :



Note that, if a student writes equations according to Perspective I, the student still has a set of equations that contradict the litmus test. But if the student then realizes that the hydrogen carbonate ion is amphoteric, another valid set of equations according to Perspective II may be written. Given the litmus test, the student should then analyze and evaluate the equations to state that Perspective II is not only correct according to modernized Arrhenius, but also consistent with the litmus test. Some students will have realized this out of part (b), and will present only Perspective II.

# Modeling Chemistry Unit 1 Answers Key

**LL Leslie**



## **Modeling Chemistry Unit 1 Answers Key:**

*Resources in Education*, 1995-10      *Chemistry in the Community*. American Chemical Society, 2002 This volume has relevance to a wide number of courses giving a hands on introduction to chemistry in relation to community issues rather than around specific chemical concepts      **Shreir's Corrosion**, 2009-02-27 This four volume reference work builds upon the success of past editions of Elsevier's Corrosion title by Shreir Jarman and Burstein covering the range of innovations and applications that have emerged in the years since its publication Developed in partnership with experts from the Corrosion and Protection Centre at the University of Manchester Shreir's Corrosion meets the research and productivity needs of engineers consultants and researchers alike Incorporates coverage of all aspects of the corrosion phenomenon from the science behind corrosion of metallic and non metallic materials in liquids and gases to the management of corrosion in specific industries and applications Features cutting edge topics such as medical applications metal matrix composites and corrosion modeling Covers the benefits and limitations of techniques from scanning probes to electrochemical noise and impedance spectroscopy      **Chemistry in the Community (Enhanced Core Four)** American Chemical Society, 2006-02-15

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*Energy Research Abstracts*, 1990 Cities and Their Vital Systems Advisory Committee on Technology and Society, 1989 Cities and Their Vital Systems asks basic questions about the longevity utility and nature of urban infrastructures analyzes how they grow interact and change and asks how when and at what cost they should be replaced Among the topics discussed are problems arising from increasing air travel and airport congestion the adequacy of water supplies and waste treatment the impact of new technologies on construction urban real estate values and the field of telematics the combination of computers and telecommunications that makes money machines and national newspapers possible

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**Energy and Water Development Appropriations for 1999** United States. Congress. House. Committee on Appropriations. Subcommittee on Energy and Water Development, 1998

**Cumulated Index Medicus**, 2000

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