At a high temperature, 0.50 mol of HBr was placed in a 1.0 L container and allowed to decompose according to the reaction:

$$2HBr(g) \rightleftharpoons H_{2(g)} + Br_{2(g)}$$

At equilibrium the [Br<sub>2</sub>] was measured to be 0.13 M. What is K<sub>eq</sub> for this reaction at this temperature?

	2HBr =	= H2 +	Br2
[x]	0.50	0	0
[C]	-0-26	+0.13	+0.13
IE]	0.24	0.13	0.13

Keg =	IH2][Br]2	= (0.13)2	=0.29
	100001100000000000000000000000000000000	150000000000000	

Answer Keg = 0.29

 When 1.0 mol of NH<sub>3(g)</sub> and 0.40 mol of N<sub>2(g)</sub> are placed in a 5.0 L vessel and allowed to reach equilibrium at a certain temperature, it is found that 0.78 mol of NH<sub>3</sub> is present. The reaction is:

2NH3(g) = 3H2(g) + N2(g)

Initial [NH3] = 1.0 mol = 0.20 M Initial [N2] = 0.40 mol = 0.080 M

Equil [NH3] = 0.78 mol = 0.156 M 2NH2 = 3 H2 + N2

	aNH3 &	= 3 Na	+ N2
[I]	0.20	0	0.080
EC3	-0.044	+0.066	+0.022
LEJ.	0.156	0.066	0.102

a) Calculate the equilibrium concentrations of all three species.

[NH<sub>3</sub>] - 0.16 M [H<sub>2</sub>] - 0.066 M [N<sub>2</sub>] - 0.10 M

b), Calculate the value of the equilibrium constant at this temperature. (use unrounded cones, then round to 2503)

 $\text{Keq} = \frac{\text{EH}_2 \text{J}^3 \text{EN}_2 \text{J}}{\text{ENH}_3 \text{J}^2} = \frac{(0.066)^3 (0.102)}{(0.156)^2} = 0.0012$ Answ

Answer 1.2×10-3

c) How many moles of H<sub>2</sub> are present at equilibrium?

0.066 M x 5.0 L = 0.33 mal

Answer 0.33 mol

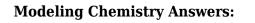
d) How many moles of N<sub>2</sub> are present at equilibrium?

0.102 M x 5.0 L = 0.51 mal

Answer 0.51 mol

# **Modeling Chemistry Answers**

**Léo Pomar** 



Unveiling the Magic of Words: A Review of "Modeling Chemistry Answers"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Modeling Chemistry Answers**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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