	100
Gas	Stoichiometry Worksheet ON Name: Name:
See T.	Volume - Volume Date: Period: Period:
1.	In the electrolysis of water, 75.0 mL of oxygen are produced. How many milliliters of hydrogen are produced? (Ans: 150. mL)
	75.0 ml Os 2 mel Hz = (150 mo Hc)
2.	How many liters of hydrogen are required to react with excess nitrogen to produce 50.0 L ammonia (NH ₂)? (Ans: 75.0 L) BH±+ W± ⇒ ≥ NH 5
	(Ans: 75.0 L) $SH_{\pm} + N_{\pm} \rightarrow 2NH_{S}$ $SO.0 \leq NH_{S} \mid 3 \leq H_{Z} = 175.0 \leq H_{Z} \mid$
3	Annmonia burns in oxygen to give nitrogen monoxide (NO) and water. If 20.0 L of ammonia are burned, how many liters of oxygen are required? (Ans: 25.0L) $20.0 \pm NH_2$ $5 \pm 0 \pm 2.0 \pm 3.0 \pm 3.0$
	Carbon monoxide burns in oxygen and forms carbon dioxide. a. How many liters of carbon dioxide are produced from 15.0 L of CO? (Ans: 15.0 L) 2202
0n5 -	b. How many liters of oxygen are needed? (Ans: 75.0 L) 15.0 £ CO 1 £ C = 2.5 £ C = 2.5 £ C = 3.5 £ C =
-	uncombined? (Ans: 200 mL)
	2H2 + 02 7 2 H2O 200NU RYCENSU 1/00 NB 400NB 7 400 NB
	400 nd 400 nd 7 400 nd
- 6	Ethylene burns in oxygen to form carbon dioxide and water vapor in the following reaction: C ₂ H _{4gt} + 3 O _{2(gt)} → 2 CO _{2(gt)} + 2 H ₂ O _(gt) . How many liters of water can be formed if 1.25 liters of ethylene are consumed in this reaction at STP? (Ans: 2.50 L)
1	25\$ C2H8 2 \$ H20 - 2.50 \$ H20 1
7075000	
III.	Mass-Volume 2 Na(s) + 2H2(1) -> 2NaOH3) + Hzg)
STY	1. How many grams of sodium are needed to release 4.0 L of hydrogen from water? (Ans: 8.21 g) 4.0 2 1/2 1.0 mol Hz 2 mol Na 23.0g Na = 8.21g Na 32.4 L Hz /mol Hz 1.0mol No 8.21g Na
	127 # 1 4 . 1 / nest 11 > 1 / UNat Vp
	2. When 130 g of zine react with excess hydrochloric acid, what volume of hydrogen is produced? (Ans: 44.8L)
31	When 130 g of zinc react with excess hydrochloric acid, what volume of hydrogen is produced? (Ans: 44.8L) $2n + 1 + 2 = 1 + 2 = 1 + 1 = 1 = 1 = 1 = 1 = 1 = 1 = 1 = 1$

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