I. ONE REACTANT IN EXCESS - COMBUSTION

You burn an excess of butane (C₄H₁₀) in a limited amount of oxygen (O₂). If 67.2 L of CO₂ gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:

STEP 2: Use a molar conversion to convert quantity of product given to moles:

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol CO}_2$$
 * (13 mol O_2) = 4.875 mol O_2
 (8 mol CO_2)

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol } O_2 \ (22.4 \text{ L} \cdot O_2) = 109.2 \text{ L} \cdot O_2 = 109 \text{ L} \cdot O_2$$

(1 mol O_2)

II. ONE REACTANT IN EXCESS - SINGLE DISPLACEMENT

14 g of lithium is combined with excess K₂SO₄. What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP I: Balance the reaction:

Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

2 mol Li *
$$(1 \text{ mol Li}_2SO_4) = 1 \text{ mol Li}_2SO_4$$

(2 mol Li)

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

Mixed Stoichiometry Problems 2 And Answer

LL Leslie

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