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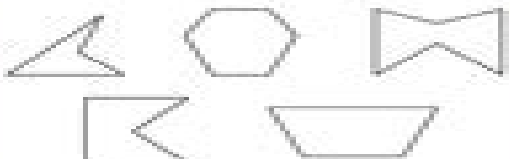
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Zhe Xu



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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CN}^{-}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{OH}^{-} + 2\text{e}^{-} \rightarrow 2\text{Au} + 2\text{H}_2\text{O} + \text{O}_2 + 2\text{OH}^{-}$ multiplied by a factor of 2 $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 + 2\text{OH}^{-} \rightarrow 2\text{H}_2\text{O}_2 + 2\text{OH}^{-}$ the final answer

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