

INORGANIC CHEMISTRY



CHEMICAL BONDING

IONIC BONDS

- Atoms with low ionization energy form cations (1-3), and atoms with high electron affinity form anions (1-3) attracted to form ionic bonds.

THE LEWIS OCTET RULE: COVALENT BONDS

- Atoms other than hydrogen (stable with 2 valence electrons) will form bonds until they are surrounded by 8 valence electrons, and have achieved stable noble gas configuration.

The Lewis Dot Diagrams consist of one dot for each valence electron surrounding the symbol of the element.



BOND DISSOCIATION ENERGIES: CALCULATING ΔH

- $\Delta H = \sum \text{Bond energies of reactants} - \sum \text{Bond energies of products}$
- Example: $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- $\Delta H = 4(413.4) + 2(495.8) - 2(463) - 2(463.0) = -890.2 \text{ kJ/mol}$ (exothermic reaction)
- Reaction Exothermic = energy (heat) flows out of system
- Endothermic = energy (heat) flows into system

Bond	Energy (kJ/mol)
H-H	436.4
H-Cl	431
H-Br	366
H-I	297
C-H	413.4
C-Cl	339
C-Br	276
C-I	213

PERCENT IONIC CHARACTER OF A CHEMICAL BOND

- Bonds with 50% or greater ionic character are ionic.
- Bonds with less than 50% ionic character are covalent.
- Transition from ionic to covalent bonding results in a reduction in electrical conductivity, melting point, and boiling point.



INTRODUCTION

- The study of compounds (group of carbon, except for CO , CO_2 , CS_2 , and compounds containing CN^- , CO_3^{2-} , and HCO_3^-)

INORGANIC COMPOUNDS

Name	Formula	Name	Formula
Alumina	Al_2O_3	Glauber's Salt	Na_2SO_4
Ammonia	NH_3	Cyanate	$\text{CaNCN} \cdot 2\text{H}_2\text{O}$
Ammonia	NH_3	Heavy Water	D_2O
Baking soda	NaHCO_3	Litharge	PbO
Bauxite	$\text{Al}_2\text{O}_3 \cdot \text{H}_2\text{O} \cdot \text{H}_2\text{O}$	Salt or Magnesia	MgO
Calcite	CaCO_3	Carbide	CaC_2
Carbon	C	Silica	SiO_2
Cyanide	CN^-	Silver Lime	CaCO_3
Dry Ice	CO_2 (solid)	Salt	NaCl
Glauber	Na_2SO_4	Whiting Soda	$\text{Na}_2\text{CO}_3 \cdot \text{H}_2\text{O}$

COMMON INORGANIC IONS

Ammonium (NH_4^+)	Mercuric (Hg_2^{2+})	Nitride (N^{3-})
Ammonium (NH_4^+)	Mercurous (Hg_2^{2+})	Hydrogen Carbonate or Bicarbonate (HCO_3^-)
Barium (Ba^{2+})	Potassium (K^+)	Hydrogen Phosphate (HPO_4^{2-})
Calcium (Ca^{2+})	Silver (Ag^+)	Hydrogen Sulfate, Bisulfate (HSO_4^-)
Cadmium (Cd^{2+})	Sodium (Na^+)	Nitrate (NO_3^-)
Chromium (Cr^{3+})	Strontium (Sr^{2+})	Fluoride (F^-)
Chromium (Cr^{6+})	Tin (Sn), Stannous (Sn^{2+})	Selenide (Se^{2-})
Cobalt (Co), Cobaltous (Co^{2+})	Zinc (Zn^{2+})	Selenate (SeO_4^{2-})
Copper (Cu), Cuprous (Cu^+)	Bromide (Br^-)	Nitrite (NO_2^-)
Copper (Cu), Cupric (Cu^{2+})	Carbonate (CO_3^{2-})	Chloride (Cl^-)
Hydrogen (H^+)	Chloride (ClO_2^-)	Perchlorate (ClO_4^-)
Iron (Fe), Ferrous (Fe^{2+})	Chloride (Cl^-)	Formate (HCO_2^-)
Iron (Fe), Ferric (Fe^{3+})	Chromate (CrO_4^{2-})	Phosphate (PO_4^{3-})
Lead (Pb), Plumbic (Pb^{2+})	Cyanide (CN^-)	Sulfate (SO_4^{2-})
Lithium (Li^+)	Dichromate ($\text{Cr}_2\text{O}_7^{2-}$)	Sulfide (S^{2-})
Magnesium (Mg), Magnesian (Mg^{2+})	Ethylhydrogen Phosphate ($\text{C}_2\text{H}_5\text{HPO}_4^-$)	Sulfate (SO_4^{2-})
Magnesium (Mg^{2+})	Fluoride (F^-)	Thiosulfate ($\text{S}_2\text{O}_3^{2-}$)

PAULI'S RULE

- Indicates that one orbital has the same energy, electrons occupy separate orbitals, and do so with opposite spins.
- Using 2p orbitals to demonstrate.

ORBITAL TYPES & SHAPES



PAULI'S EXCLUSION PRINCIPLE

- No 2 electrons in an atom can have the same 4 quantum numbers. No more than 2 electrons can occupy any 1 orbital.

Inorganic Elements Student Guide

**John Krenos, Joseph Potenza, Loretta
Jones, Lynn Koplitz, Thomas Spence**



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Table of Contents Inorganic Elements Student Guide

1. Understanding the eBook Inorganic Elements Student Guide
 - The Rise of Digital Reading Inorganic Elements Student Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Inorganic Elements Student Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Inorganic Elements Student Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Inorganic Elements Student Guide
 - Personalized Recommendations
 - Inorganic Elements Student Guide User Reviews and Ratings
 - Inorganic Elements Student Guide and Bestseller Lists

5. Accessing Inorganic Elements Student Guide Free and Paid eBooks
 - Inorganic Elements Student Guide Public Domain eBooks
 - Inorganic Elements Student Guide eBook Subscription Services
 - Inorganic Elements Student Guide Budget-Friendly Options
6. Navigating Inorganic Elements Student Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Inorganic Elements Student Guide Compatibility with Devices
 - Inorganic Elements Student Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Inorganic Elements Student Guide
 - Highlighting and Note-Taking Inorganic Elements Student Guide
 - Interactive Elements Inorganic Elements Student Guide
8. Staying Engaged with Inorganic Elements Student Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Inorganic Elements Student Guide
9. Balancing eBooks and Physical Books Inorganic Elements Student Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Inorganic Elements Student Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Inorganic Elements Student Guide
 - Setting Reading Goals Inorganic Elements Student Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Inorganic Elements Student Guide
 - Fact-Checking eBook Content of Inorganic Elements Student Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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