

## Determining the Mole Ratios in a Chemical Reaction

A balanced chemical reaction equation gives the mole ratios of the reactants and the products as coefficients. When some of the chemical formulas are not known, an experiment must be conducted to help determine the mole ratios.

This experiment uses two substances, an unidentified acid and sodium hydroxide ( $\text{NaOH}$ ), which react according to the unbalanced and incomplete reaction equation below.



It is possible to identify the coefficients,  $a$  and  $b$ , for the reactants, without knowing the identity of the acid and the products of the reaction. The process that you will use to determine the coefficients is called *continuous variations*. You will prepare a series of mixtures of the two reactants. Each mixture will have the same total volume and the same total number of moles of reactants. The reaction is exothermic, thus the mixture that generates the most heat energy will be the reaction that completely consumes both the unidentified acid and the sodium hydroxide. By analyzing the results of the series of reactions, you will determine the mole ratio of the reactants.

You will also be determining whether the unidentified acid is monoprotic, diprotic, or triprotic. *Monoprotic acids* are able to donate one proton per molecule. Common examples of monoprotic acids are hydrochloric acid ( $\text{HCl}$ ) and acetic acid ( $\text{HC}_2\text{H}_3\text{O}_2$ ). *Polyprotic acids*, in contrast, can donate more than one proton per molecule. Sulfuric acid ( $\text{H}_2\text{SO}_4$ ) and carbonic acid ( $\text{H}_2\text{CO}_3$ ) are common diprotic acids that are able to donate two protons per molecule. Phosphoric acid ( $\text{H}_3\text{PO}_4$ ) and citric acid ( $\text{H}_3\text{C}_6\text{H}_7\text{O}_7$ ) are common triprotic acids.

### OBJECTIVES

- Measure the temperature changes of a series of reactions.
- Determine the stoichiometry of an oxidation-reduction reaction in which the reactants are known but the products are unknown.
- Determine whether the unidentified acid is monoprotic, diprotic, or triprotic.

# Hspi The Pogil Project Limiting Reactants

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## **Table of Contents Hspi The Pogil Project Limiting Reactants**

1. Understanding the eBook Hspi The Pogil Project Limiting Reactants
  - The Rise of Digital Reading Hspi The Pogil Project Limiting Reactants
  - Advantages of eBooks Over Traditional Books
2. Identifying Hspi The Pogil Project Limiting Reactants
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hspi The Pogil Project Limiting Reactants
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hspi The Pogil Project Limiting Reactants
  - Personalized Recommendations
  - Hspi The Pogil Project Limiting Reactants User Reviews and Ratings
  - Hspi The Pogil Project Limiting Reactants and Bestseller Lists
5. Accessing Hspi The Pogil Project Limiting Reactants Free and Paid eBooks
  - Hspi The Pogil Project Limiting Reactants Public Domain eBooks
  - Hspi The Pogil Project Limiting Reactants eBook Subscription Services
  - Hspi The Pogil Project Limiting Reactants Budget-Friendly Options
6. Navigating Hspi The Pogil Project Limiting Reactants eBook Formats

- 
- ePub, PDF, MOBI, and More
  - Hspi The Pogil Project Limiting Reactants Compatibility with Devices
  - Hspi The Pogil Project Limiting Reactants Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Hspi The Pogil Project Limiting Reactants
    - Highlighting and Note-Taking Hspi The Pogil Project Limiting Reactants
    - Interactive Elements Hspi The Pogil Project Limiting Reactants
  8. Staying Engaged with Hspi The Pogil Project Limiting Reactants
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Hspi The Pogil Project Limiting Reactants
  9. Balancing eBooks and Physical Books Hspi The Pogil Project Limiting Reactants
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Hspi The Pogil Project Limiting Reactants
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Hspi The Pogil Project Limiting Reactants
    - Setting Reading Goals Hspi The Pogil Project Limiting Reactants
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Hspi The Pogil Project Limiting Reactants
    - Fact-Checking eBook Content of Hspi The Pogil Project Limiting Reactants
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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