

Equilibria Quiz SCH4U #11 Name: _____

 K T C Assessment of Learning (Product)

10 10 5

Multiple Choice: Circle the most correct answer. (10K)

1. For the endothermic reaction: $\text{CaCO}_{3(s)} \rightleftharpoons \text{CaO}_{(s)} + \text{CO}_{2(g)}$, Le Chatelier's principle predicts that _____ will result in an increase in the number of moles of CO_2 .

- A) increasing the temperature B) decreasing the temperature
C) increasing the pressure D) removing some of the $\text{CaCO}_{3(s)}$
E) none of the above

2. Consider the following reaction at equilibrium: $2\text{NH}_{3(g)} \rightleftharpoons \text{N}_{2(g)} + 3\text{H}_{2(g)}$. Le Chatelier's principle predicts that the moles of $\text{H}_{2(g)}$ in the reaction container will increase with _____.

- A) some removal of $\text{NH}_{3(g)}$ from the reaction vessel (V and T constant)
B) a decrease in the total pressure (T constant)
C) addition of some $\text{N}_{2(g)}$ to the reaction vessel (V and T constant)
D) a decrease in the total volume of the reaction vessel (T constant)
E) an increase in total pressure by the addition of helium gas (V and T constant)

3. The value for K_{eq} for a reaction is found to be 3.2×10^{-10} . What does this imply?

- A) products are favoured B) reactants are favoured
C) reactant concentration = product concentration D) the system was open

4. Which of the following does **not** apply to all chemical equilibrium systems?

- A) They are closed.
B) The macroscopic properties are constant.
C) Forward and reverse reaction rates are equal.
D) There are equal concentrations of reactants and products.

5. What is the equilibrium constant expression for the reaction: $\text{CaCO}_3(s) \rightleftharpoons \text{CaO}(s) + \text{CO}_2(g)$?

$$A) K = [\text{CO}_2] \qquad B) K = \frac{[\text{CaO}]}{[\text{CaCO}_3]} \qquad C) K = \frac{[\text{CaCO}_3]}{[\text{CaO}][\text{CO}_2]} \qquad D) K = \frac{[\text{CaO}][\text{CO}_2]}{[\text{CaCO}_3]}$$

6. The reaction has the K shown: $2\text{N}_2\text{O}(g) + 3\text{O}_2(g) \rightleftharpoons 2\text{N}_2\text{O}_5(g)$ $K = 2.9 \times 10^4$. If a reaction mixture has the concentrations as shown, in what direction is the reaction proceeding?

- Concentrations: $[\text{N}_2\text{O}] = 1.32 \times 10^{-4} \text{ M}$ $[\text{O}_2] = 2.24 \times 10^{-4} \text{ M}$ $[\text{N}_2\text{O}_5] = 1.20 \times 10^{-4} \text{ M}$
A) Right to left B) Left to Right

Holt Chemistry Quiz Answer Key Chemical Equilibrium

J Dewey



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