

Self-Assessment Paper

Topic: Describe your personality at work, using either the Meyers-Briggs typology, or other widely recognized personality characteristics. Give specific suggestions for pitfalls to be aware of and adaptations you may have to make, in order to improve your effectiveness.

Understanding who I am is critical in order to improve myself and to be successful in the workplace. In this paper, I will analyze my personality at work based on my managerial experience and the result of the Meyers-Briggs typology. From the assessment of my personality at work, I will reflect upon possible solutions for my weaknesses in order to improve my effectiveness in the workplace.

For three years, I was an academic supervisor at Young-do English Institution, which helps K-12 students learn the language from English-native instructors in South Korea. I was in charge of training and managing foreign teachers and counseling students and parents. Since most of my duties were related to dealing with different individuals and groups of people, I could discover my strong and weak points while interacting with people at my job.

According to Meyers-Briggs typology, my personality type is SP, so one of my strong points is negotiating. Fortunately, major responsibilities of my job included negotiating and communicating among different parties. For instance, when Korean executives released new policies, I had to convey the announcement to our faculty members and had to persuade why teachers should follow this direction. Since the culture and the way of thinking are different from

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Albert A Gayle



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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq} + \text{CrO}_4^{2-} \text{aq} \rightarrow \text{Ag}_2\text{CrO}_4 \text{s}$

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web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3 \text{ aq} + \text{KCl aq} \rightarrow \text{AgCl s} + \text{KNO}_3 \text{ aq}$ total ionic $\text{Ag}^+ \text{ aq} + \text{NO}_3^- \text{ aq} + \text{K}^+ \text{ aq} + \text{Cl}^- \text{ aq} \rightarrow \text{AgCl s} + \text{K}^+ \text{ aq} + \text{NO}_3^- \text{ aq}$ net ionic $\text{Ag}^+ \text{ aq} + \text{Cl}^- \text{ aq} \rightarrow \text{AgCl s}$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{ aq} + \text{OH}^- \text{ aq} \rightarrow \text{H}_2\text{O l}$ the spectator ions in KOH aq and $\text{HNO}_3 \text{ aq}$ are $\text{K}^+ \text{ aq}$ and $\text{NO}_3^- \text{ aq}$ how to identify and list the

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web the $\text{Ag}^+ \text{ aq}$ and $\text{Cl}^- \text{ aq}$ ions become AgCl s but the $\text{Na}^+ \text{ aq}$ ions and the $\text{NO}_3^- \text{ aq}$ ions stay as $\text{Na}^+ \text{ aq}$ ions and $\text{NO}_3^- \text{ aq}$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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