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For the chemical reaction

Chemical	State	Stoichiometric Coefficient	Standard Enthalpy of Formation, $\Delta H_f^\circ$ (kJ/mol)	Standard Enthalpy of Combustion, $\Delta H_c^\circ$ (kJ/mol)
$\text{H}_2$	(g)	2	0	0
$\text{O}_2$	(g)	1	0	0
$\text{H}_2\text{O}$	(l)	2	-285.8	-285.8

The grams of water that can be produced from 2.192g of hydrogen can be calculated as follows.

$(2.192 \text{ g H}_2) \left( \frac{1 \text{ mol H}_2}{2.016 \text{ g H}_2} \right) \left( \frac{2 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2} \right) \left( \frac{18.015 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \right)$

$(2.192 \text{ g}) \left( \frac{1 \text{ mol H}_2}{2.016 \text{ g}} \right)$

# Guide To Microsoft Equation Editor 3.0

**Karin Nielsen-Saines**



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