

# Stoichiometry - Gram Conversions

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1. Use the following balanced equation to answer the questions below:



- How many grams of  $\text{O}_2$  are produced when 34.5g of  $\text{HgO}$  decompose?
  - How many grams of  $\text{HgO}$  need to decompose to produce 0.277g of  $\text{Hg}$ ?
  - If 52.1g of  $\text{HgO}$  decomposes, then how many grams of  $\text{Hg}$  will be made?
2. When catalyzed by sulfuric acid, potassium chlorate and sucrose react to produce heat,  $\text{CO}_2$ ,  $\text{H}_2\text{O}$ , and  $\text{KCl}$ . Below is a balanced equation for this reaction.



- How many grams of  $\text{CO}_2$  will form if 52.9g of sucrose react completely with  $\text{KClO}_3$ ?
  - How many grams of  $\text{KClO}_3$  are needed to completely react with 781g of  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ ?
  - How many moles of water will be produced if 99.1g of  $\text{KClO}_3$  reacts completely with sucrose?
3. Nitrogen dioxide reacts with water to produce nitric acid and nitric oxide. Below is the balanced equation for this reaction.



- How many grams of  $\text{NO}$  will be made if 5.57g of  $\text{NO}_2$  reacts completely with water?

# Gram To Gram Stoichiometry Worksheet Answers

**M Walker**



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