

Residential Gas Water Heater **USE & CARE MANUAL**

WITH INSTALLATION INSTRUCTIONS FOR THE CONTRACTOR



The purpose of this manual is twofold: for the installing contractor, to provide requirements and recommendations for the proper installation and adjustment of the water heater, and for the owner-operator, to explain the features, operation, safety precautions, maintenance and trouble shooting of the water heater. This manual also includes a parts list.

It is imperative that all persons who are expected to install, operate or adjust this water heater read the instructions carefully so that they may understand how to do so.

Any questions regarding the operation, maintenance, service or warranty of this water heater should be directed to the entity from whom it was purchased. If additional information is required, refer to the section on How to Obtain Service Assistance.

Do Not Destroy this Manual. Please read carefully and keep in a safe place for Future Reference.



Recognize this symbol as an indication of important Safety Information!



WARNING: If the information in these instructions are not followed exactly, a fire or explosion may result causing property damage, personal injury or death.



FOR YOUR SAFETY!

- Do not store or use gasoline or other flammable vapors or liquids or other combustible materials in the vicinity of this or any other appliance. To do so may result in an explosion or fire.

— WHAT TO DO IF YOU SMELL GAS

- Do not try to light any appliance.
- Do not touch any electrical switch; do not use any phone in your building.

- Immediately call your gas supplier from a neighbor's phone. Follow the gas supplier's instructions.
 - If you cannot reach your gas supplier, call the fire department.
 - Do not return to your home until authorized by the gas supplier or fire department.
- Improper installation, adjustment, alteration, service or maintenance can cause injury, property damage or death. Refer to this manual. Installation and service must be performed by a qualified installer, service agency or the gas supplier.

Ge Gas Heater User Manual

AW Chickering

A red circular graphic with a gradient, appearing as a stylized arrow or a light effect, positioned to the right of the 'AW Chickering' text.

Ge Gas Heater User Manual:

The Enigmatic Realm of **Ge Gas Heater User Manual**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Ge Gas Heater User Manual** a literary masterpiece penned by way of a renowned author, readers set about a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting affect the hearts and minds of those that partake in its reading experience.

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this video accompanies the chm 252 lab nitration of bromobenzene in this video students will get an introduction to the lab as well as the set up for conducting the reaction students should

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction e.g. particularly nitration of methyl benzoate to give methyl m-nitrobenzoate thus determine the weight, melting point and percentage yield of the pure methyl m-nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene. Bromine substituted on an aromatic ring is an ortho/para directing group. It will direct an incoming electrophile to the ortho and para positions. It is also deactivating, meaning that its reactions will be slower than the identical reactions of benzene.

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Section 20711 Lab Experiment: Nitration of Bromobenzene. Introduction: The chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule. Nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid. This mixture is useful to obtain the active nitronium ion.

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1. Go to the restroom and wash the affected area with lots of soap and water. The soap in the lab is not suitable for this purpose.
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1.

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In the nitration of bromobenzene, a mixture of sulfuric acid and nitric acid reacts with one another. Nitric acid, the weaker of the two acids, will serve as the base to form the nitronium ion. Bromobenzene will attack the nitronium ion by nucleophilic attack.

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 °C

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture HNO_3 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by the nitration of bromobenzene

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lab report 15 nitration of bromobenzene brynn abbott amp katie muerer bea0006 uah ch 336 02 september 21 2022 abstract in this experiment the electrophilic aromatic substitution mechanism was used to obtain 4-nitrobromobenzene and 2-nitrobromobenzene recrystallization was used to purify 4-nitrobromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic

substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1 h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

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chem 2261 lab 5 nitration of bromobenzene introduction an electrophilic aromatic substitution reaction is an important process that allows for intermediate chemical structures to be used as precursors to pharmaceuticals and other industrially produced chemicals by adding different functional groups into an aromatic ring

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lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid creates what is known as a nitro group

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