

Unit 1 Test STANDARD LEVEL

NAME _____
CLASS _____

A2

Grammar

- Complete the sentences with the correct present simple form of the verbs given.
 - We _____ (not go) to school on Saturdays.
It's closed.
 - _____ Anna _____ (like) art classes?
She has lots of coloured pencils.
 - My dad _____ (not work) near our house. He drives a long way to get there.
 - A: Do you get up early on Sundays?
B: No, I _____ (do).
 - John _____ (play) football after school.

/ 5
 - Complete the sentences with *at*, *in* or *on*.
 - We have a morning break _____ 11.15.
 - We don't have afternoon classes _____ Wednesdays.
 - Can you come to my party? It's _____ the sixth of May.
 - I like playing on my computer _____ the evening.
 - It's very cold here _____ January.

/ 5
 - Complete the sentences with the correct answer, **A** or **B**.
 - We _____ have biology on Mondays. It's my favourite lesson so I'm never late.
A - always
B - hardly ever
 - My mum _____ drives me to school, but only when the bus is late.
A - never
B - sometimes
 - Brett _____ cleans his dad's car at the weekend. I see him do it nearly every Saturday.
A - often
B - hardly ever
 - My football practice _____ starts at a quarter to four but today it's at half past four.
A - never
B - usually
 - Our teacher _____ gives us homework at the weekend. She says it's time to relax.
A - never
B - often

/ 5
- Total / 15**

Vocabulary

- Choose the correct alternatives to complete the sentences.
 - I ~~do make~~ my homework in my bedroom before dinner.
 - My sister ~~does has~~ lunch in the school cafeteria.
 - We study ~~geography/physics~~ on Thursdays. I love learning about different countries.
 - John is very good at ~~art/math~~. He likes solving problems with numbers.
 - I usually ~~finish up~~ to bed when I'm tired.

/ 5
 - Complete the sentences with words for places to go in a town.
 - There's a really nice _____ near my school. Sometimes we go swimming on the way home.
 - When I want to read a book, I don't buy one. I go to the _____ instead.
 - My dad takes me to the _____ when it rains. We like learning about the history of our town.
 - The _____ near our house is very big. It has lots of shops.
 - The _____ in our town is fantastic. We often have picnics there in summer.

/ 5
 - Complete the sentences with the correct form of these verbs. There are two extra options.

go - listen - make - play - surf - take
watch

 - She _____ shopping at the weekend.
 - I _____ the internet in the evening.
 - He _____ the guitar.
 - They _____ films on their computer.
 - I don't like _____ photos of people.

/ 5
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John R. Vacca



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web acids and bases neutralisation reactions in an acid base titration the concentration of the acid or base is determined by accurately measuring the volumes used in the neutralisation reaction an indicator can be added to show the end point of the reaction chemistry in society chemical analysis analytical methods

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web introduction this laboratory exercise relies on a titration technique to determine an unknown concentration of

monoprotic acid in solution in the process of titration a basic solution is gradually added to the acidic solution until complete neutralization is obtained

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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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web lab report on acid base titration 1346 words 6 pages titration process is used in an acid base experiment in order to determine the concentrations of solutions of acids and bases through the titration process we are able to identify physical changes to the mixture such as the colour change to indicate the end point of the experiment

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web aug 30 2022 acid base titrations are usually used to find the amount of a known acidic or basic substance through acid base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known molarity that will react with the analyte

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web chem 1210 spring 2019 experiment 10.11 part 1 acid base titration abstract the purpose of this experiment is to observe the titration of hydrochloric acid a strong acid with sodium hydroxide a strong base and acetic acid a

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1

$\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{O l}$ one mole of hcl would be fully neutralized by one mole of naoh

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide naoh and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

lab 6 acid base titration chemistry libretexts - Feb 10 2023

web jul 19 2023 introduction titration is an analytical quantitative technique used to determine the concentration of a solute a pH titration is used to determine the concentration of an acid or a base titrations play an important role in determining amount and purity in many manufacturing processes

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

lab report titration chm1046l volumetric analysis acid base - Apr 12 2023

web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 2\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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data calculations and conclusion acid base titration lab - Jun 02 2022

web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2M HCl with the NaOH of unknown concentration in order to experimentally

ascertain the concentration of the NaOH

this is a chemistry lab report on an acid base titration experiment - Dec 28 2021

web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction

$\text{NaOH (aq)} + \text{C}_6\text{H}_4\text{COOH (aq)} \rightarrow \text{C}_6\text{H}_4\text{COONa (aq)} + \text{H}_2\text{O (l)}$ the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the