

GAS LAWS

AS EASY AS ABCD!

Avogadro's Law:

$$V = an, \text{ or } \frac{V_1}{n_1} = \frac{V_2}{n_2}$$

Boyle's Law:

$$PV = k, \text{ or } P_1V_1 = P_2V_2$$

Charles' Law:

$$V = bT, \text{ or } \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Dalton's Law:

$$P_{\text{TOTAL}} = P_1 + P_2 + P_3 \dots$$

IDEAL GAS LAW:

$$PV = nRT$$

$$R = \frac{0.08206 \text{ L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$$

Conversion
Factors

1 atm

= 760 mm Hg

= 760 torr

= 101325 Pa

= 101.325 kPa

= 1.01325 bar

= 29.92 in Hg

= 14.7 lb/in²

= 101325 N/m²

= pressure at
sea level

Gas Laws Guide

Pamela Mills, Amina El-Ashmawy



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