			Fig.

K.E.Y.

Period Student Number

Gas Law Stoichiometry Worksheet

SHOW	Directions: Use significant figures and units in the problems below.

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:

2 Na₃₀ +

2 NaClos

 What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal,

CaH_B(a)

2.34 & Cl2

H H2Om

a. Balance the above reaction.

 What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

m= X

V = U2(X)

T1 = 25+273 = 298K

P1 = 1.05 adm. V = (0.628 n +00821 de oft 2816) 44 of 1001

1.05 adm. V = (0.628 n +00821 de oft 2816) 44 of 1001

C3HK

3. Potassium permanganate, KMnO₄, is produced commercially by oxidizing aqueous potassium manganate, K2MnO4.

2 KoMnO4 (a) +

Clarge

2 KMnO_{4 (a)}

a KCla

Balance the above reaction.

What volume of Cl₂(g), measured at STP, is needed to produce 10.0 g of

KMinO₄?

10.0g Km, 04 10012 1 0.709 L nitride, ammonia cas in 22 2011 1001612 0.709 L

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated 3 H₂O_(m) 3 MgO_(m) xnort Michigan

Balance the above reaction.

PV=mRT b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of

10.32 My 3 N2 | 1 more 2 NH = 0.204 mol 294+273 = $\frac{5.03}{1521760}$ V = $\frac{100.92}{1521760}$ | 1 more 2 NH = 0.204 mol 297 K T521760 V = $\frac{100.92}{1521760}$ | 10.33 H20 | 1 mol 2 mol 2 mol 2989 M = 0.989 M = 0.989

Gas Stoichiometry Answer Key

Xiaolong Qi

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