

Name: Emarie Johnson

Gas Laws Webquest

1. What is a **gas**? Something that is in the air.
2. What does the **kinetic molecular theory** explain? (at least 3 things) the particles in a gas are in constant, random motion, the combined volume of the particles is negligible, the particles exert no forces on one another
3. Describe the three states of matter in terms of how they fill a container. Solid is the state in which matter maintains a fixed volume and shape, liquid is the state in which matter adapts to the shape of its container but varies only slightly in volume, and gas is the state in which matter expands to occupy the volume and shape of its container.
4. Describe the typical motions of atoms in a solid, a liquid and a gas. gas vibrate and move freely at high speeds. liquid vibrate, move about, and slide past each other. solid vibrate but generally do not move from place to place.
5. What are the 4 physical (measurable) characteristics of a gas?
Mass, Volume, Temperature, Pressure
6. The Kelvin scale is based on the concept of absolute zero. What is absolute zero, and what happens to particles at absolute zero? the particles stop moving and all disorder disappears.

Comparison of Temperature Scales

	Fahrenheit	Celsius	Kelvin
Water boils	212	100	373.1 k
Body temperature	97	37	310.2
Water freezes	32	0	273 k
Absolute zero	-459.67	-273.15	0 k

7. What are the boiling points of Oxygen, Nitrogen and Fluorine gas? -183, -196, -188
8. Who invented the mercury barometer and what is it used to measure? Evangelista Torricelli
9. How does the barometer work? Insert a quick labeled diagram. The needle turns clockwise when the pressure is rising
10. What causes the pressure that a gas exerts on the walls of its container? The moving particles in a gas collide with each other and also with the walls of the container
11. Why do we not feel the weight of the atmosphere? Because air is a fluid, the weight of the air is transmitted to the palm of your hand and to the back of your hand at the same time.
12. What happens to the pressure of a confined gas at a constant temperature when the volume is reduced by 1/2? increase to twice its original value.
13. What happens to the volume of a gas at constant temperature when the pressure is increased? inversely proportional to the pressure
14. Write Boyle's law using both words and a formula. $P_1V_1 = P_2V_2$

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Didier Musso



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