Name:	KEY	Block:

Assignment #2 - Compound Names and Formulas Multi-valent ions only

A. Write the correct formula for the	e following compounds, all of which have been named using the
modern Roman Numeral Method.	The ionic charge is given after the first element (metallic) in Roman
Numerals.	
~ ~	

1. copper (II) oxide	11. manganese (III) oxide Mn 203
2. mercury (I) oxide	12. vanadium (II) bromide VB(z
3. gold (III) chloride _AuCl3	13. niobium(V) oxide No205
4. thallium (III) bromide TUSV 3	14. titanium (III) oxide T\20>
5. bismuth (V) oxide <u>B1205</u>	15. titanium (III) nitride TIN
6. terbium (III) oxide	16. iron (II) oxide <u>Fe.O</u>
7. uranium (VI) oxide \(\lambda \)	17. cobalt (II) phosphide C03P2
8. protactinium (IV) bromide PhBY4	18. tin (II) oxide
9. cerium (III) oxide	19. thulium (II) bromide TWBY 2
10. arsenic (V) sulphideAS2S5	20. copper (II) bromide CURY2
B. Name the compound using the Roman Numer	
1. SnCl Tin (IV) Chloride	9. PdF. Palladium (IV) Fluoride
2. BIBIS BISMINTO BIOMICE	10. Os2Os OSMIUM (III) Oxide
3. POO2 POLONIUM (IV) OXICE	11. MOBY MOLYOCHNIAM (II) Bramide
4. Pb/2 100d(II) Iodide	12. VOIS Vanadium (V) Chloride
5. Hgo Mercury (II) Oxide	13. Mn ₂ O ₃ Managamese (III) Oxide
6. HgCI _ MEYCUYY (I) Chroride_	
7. Augos Gold (III) Oxide	15. NP2O3 NEDTURNIUM (II) OXKIC
8. FeCla Iron (II) Chlorida	16. V2Os Vanadium (V) Oxide

Formula Of Ionic Compound Lab Answer Key

RJ Alexander

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