

Exercise 14.90 with feedback

You may want to reference ([10 pages 535 - 543](#)) section 14.7 while completing this problem.

A 0.10 m aqueous solution of an ionic compound with the formula MX has a freezing point of $-2.3\text{ }^{\circ}\text{C}$.

Part A

Calculate the van't Hoff factor (i) for MX at this concentration.

Express your answer using two significant figures.

$i =$

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Mastering Chemistry Answer Key Chapter 1

Wolfgang Guggemos



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In a global consumed by screens and the ceaseless chatter of instantaneous conversation, the melodic beauty and emotional symphony produced by the written word usually fade into the backdrop, eclipsed by the constant noise and distractions that permeate our lives. But, located within the pages of **Mastering Chemistry Answer Key Chapter 1** an enchanting fictional prize overflowing with fresh emotions, lies an immersive symphony waiting to be embraced. Crafted by an outstanding composer of language, this captivating masterpiece conducts viewers on a mental trip, well unraveling the hidden songs and profound affect resonating within each carefully constructed phrase. Within the depths of the emotional assessment, we will discover the book is key harmonies, analyze their enthralling publishing type, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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