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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} + 2\text{Ag}^+ \text{aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{Ag}(\text{aq}) + 2\text{F}(\text{aq}) + 2\text{NH}_4(\text{aq}) + \text{Cr}_2\text{O}_7^{2-}(\text{aq}) \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7(\text{s}) + 2\text{NH}_4(\text{aq}) + 2\text{F}(\text{aq})$ because two $\text{NH}_4(\text{aq})$

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