

Key

## Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant  $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by  $101.3 \text{ kPa} / 1 \text{ atm}$  to get  
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of  $87^\circ\text{C}$ , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

$$\text{or } P = 166.29 \text{ kPa}$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of  $56^\circ\text{C}$ , what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 327 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of  $67^\circ\text{C}$ , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 390 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

$$\text{or } P = 540.61 \text{ kPa}$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

# Gases Properties And Laws Answers

**Xiang Xie**



## **Gases Properties And Laws Answers:**

## Decoding **Gases Properties And Laws Answers**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Gases Properties And Laws Answers**," a mesmerizing literary creation penned with a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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