

Name

Kary

Date

Period

## GAS STOICHIOMETRY WORKSHEET

Please answer the following using proper units and showing all dimensional analysis. Please note that these problems require a balanced chemical equation.

1. Carbon monoxide reacts with oxygen to produce carbon dioxide. Answer the following questions for the reaction of 1.0 L of carbon monoxide and oxygen at STP.

- a. Write the balanced equation.



- b. How many liters of oxygen are required to react?

$$1.0\text{ L CO} \times \frac{1\text{ mol CO}}{22.4\text{ L}} \times \frac{1\text{ mol O}_2}{2\text{ mol CO}} \times \frac{22.4\text{ L}}{1\text{ mol}} = 0.50\text{ L O}_2$$

- c. How many liters of carbon dioxide are produced?

$$1.0\text{ L CO} \times \frac{1\text{ mol CO}}{22.4\text{ L}} \times \frac{2\text{ mol CO}_2}{2\text{ mol CO}} \times \frac{22.4\text{ L CO}_2}{1\text{ mol CO}_2} = 1.00\text{ L CO}_2$$

2. Acetylene gas ( $\text{C}_2\text{H}_2$ ) undergoes combustion to produce carbon dioxide and water vapor.

- a. Write the balanced equation.



- b. How many liters of  $\text{C}_2\text{H}_2$  are required to produce 75.0 L of  $\text{CO}_2$ ?

$$75.0\text{ L CO}_2 \times \frac{1\text{ mol CO}_2}{22.4\text{ L CO}_2} \times \frac{2\text{ mol C}_2\text{H}_2}{4\text{ mol CO}_2} \times \frac{22.4\text{ L C}_2\text{H}_2}{1\text{ mol C}_2\text{H}_2} = 37.50\text{ L C}_2\text{H}_2$$

- c. What volume of  $\text{H}_2\text{O}$  is produced?

$$75.0\text{ L CO}_2 \times \frac{1\text{ mol CO}_2}{22.4\text{ L}} \times \frac{2\text{ mol H}_2\text{O}}{4\text{ mol CO}_2} \times \frac{22.4\text{ L H}_2\text{O}}{1\text{ mol H}_2\text{O}} = 37.50\text{ L H}_2\text{O}$$

- d. What volume of  $\text{O}_2$  is required?

$$75.00\text{ L CO}_2 \times \frac{1\text{ mol CO}_2}{22.4\text{ L CO}_2} \times \frac{5\text{ mol O}_2}{4\text{ mol CO}_2} \times \frac{22.4\text{ L O}_2}{1\text{ mol O}_2} = 93.75\text{ L O}_2$$

3. If liquid carbon disulfide ( $\text{CS}_2$ ) reacts with 450 mL of oxygen to produce the gases carbon dioxide and sulfur dioxide, what volume of each product is produced?  $450\text{ mL} = .45\text{ L}$

- a. Write the balanced equation?



- b. What volume of carbon dioxide is produced?

$$.45\text{ L O}_2 \times \frac{1\text{ mol O}_2}{22.4\text{ L}} \times \frac{1\text{ mol CO}_2}{3\text{ mol O}_2} \times \frac{22.4\text{ L CO}_2}{1\text{ mol CO}_2} = 0.15\text{ L CO}_2$$

- c. What volume of sulfur dioxide is produced?

$$.45\text{ L O}_2 \times \frac{1\text{ mol O}_2}{22.4\text{ L}} \times \frac{2\text{ mol SO}_2}{3\text{ mol O}_2} \times \frac{22.4\text{ L SO}_2}{1\text{ mol SO}_2} = 0.30\text{ L SO}_2$$

# Gas Stoichiometry Answer Key 13

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