Gas Laws: Combined Gas Law

The following problems may be solved using the relationships described by either Charles' or Boyles' laws, together in the combined gas law. Show all the steps to the solution of the problem.

 $V_1P_1/T_1 = V_2P_2/T_2$

 Carbon dioxide occupies a 2.54 dm³ container at STP. What will be the volume when the pressure is 150 KPa and 26°C?

 $V_3 = [2.54 \text{ dm}^3][101.3 \text{ KPa}][26^{\circ}\text{C} + 273\text{K}] = 1.88 \text{ dm}^3$ [150 KPa][273 K]

Oxygen occupies a fixed container of 5.5L at STP. What will happen to the pressure if the temperature rises to 300K?

P₂ = [5.5 L] [101.3 KPa][300 K] = 111.3 KPa [5.5 L] [273 K]

3. Methane is compressed in a closed 15.8 dm³ container at 101.3 KPa. If the volume drops to 8.7 dm³ and the temperature begins at 25°C and then drops to 18°C, what will the pressure of the gas be?

P₂ = [101.3 KPa][15.8 dm³][300 K] = 179.6 KPa [8.7 dm³][273 K]

4. A helium balloon is fully inflated at 1.2 L. When the clerk is filling the balloon, she stops to make sure it is not going to explode and checks the pressure. The pressure is 5.7 KPa and the temperature at the store is 24°C. The balloon is 0.75L full. She stops when the balloon is fully inflated and the pressure is 7.2 KPa. When the customer takes the balloon outside, it explodes. What was the temperature outdoors?

 $T = [24^{\circ}C + 273K][1.2 L][7.2 KPa] = 600 K$ [0.75 L][5.7 KPa]

5. Oxygen gas is added to a rebreather at the same rate as it would be in the open air. The volume is 4.8 dm³ and the temperature is 25 °C. The pressure is1 atm. Once the tank is lowered under the water the pressure increases to 2 atm and the temperature drops to 12°C. Is the tank in danger of exploding if it's maximum volume is 5.0 dm³?

 $V_2 = [4.8 \text{ dm}^3][1 \text{ atm}][12^{\circ}C + 273 \text{K}] = 2.1 \text{ dm}^3$, not in danger of exploding $[2 \text{ atm}][25^{\circ}C + 273 \text{K}]$

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which then join briefly to form O₃, ozone. Your olfactory sense can detect 1.00 mL
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- 1. Understanding the eBook Gas Laws University Of Virginia Scuba Answers
 - o The Rise of Digital Reading Gas Laws University Of Virginia Scuba Answers
 - o Advantages of eBooks Over Traditional Books
- 2. Identifying Gas Laws University Of Virginia Scuba Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws University Of Virginia Scuba Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Gas Laws University Of Virginia Scuba Answers

- Personalized Recommendations
- Gas Laws University Of Virginia Scuba Answers User Reviews and Ratings
- Gas Laws University Of Virginia Scuba Answers and Bestseller Lists
- 5. Accessing Gas Laws University Of Virginia Scuba Answers Free and Paid eBooks
 - Gas Laws University Of Virginia Scuba Answers Public Domain eBooks
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 - Gas Laws University Of Virginia Scuba Answers Budget-Friendly Options
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 - o ePub, PDF, MOBI, and More
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 - Gas Laws University Of Virginia Scuba Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Gas Laws University Of Virginia Scuba Answers
 - Interactive Elements Gas Laws University Of Virginia Scuba Answers
- 8. Staying Engaged with Gas Laws University Of Virginia Scuba Answers
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws University Of Virginia Scuba Answers
- 9. Balancing eBooks and Physical Books Gas Laws University Of Virginia Scuba Answers
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Gas Laws University Of Virginia Scuba Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Laws University Of Virginia Scuba Answers
 - Setting Reading Goals Gas Laws University Of Virginia Scuba Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Laws University Of Virginia Scuba Answers

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- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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