

12 • The Gas Laws**THE COMBINED GAS LAW**

In practical terms, it is often difficult to hold any of the variables constant. When there is a change in pressure, volume and temperature, the combined gas law is used.

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \text{or} \quad P_1 \times V_1 \times T_2 = P_2 \times V_2 \times T_1$$

$$K = ^\circ C + 273$$

*See
Scratch
paper*

Complete the following chart.

	P_1	V_1	T_1	P_2	V_2	T_2
1	1.50 atm	3.00 L	20.0°C 293 K	2.50 atm		30.0°C 303 K
2	720. torr	256. mL	25.0°C 298 K		250. mL	50.0°C 323 K
3	600. mmHg	2.50 L	22.0°C 325 K	760. mmHg	1.80 L	
4		750. mL	0.00°C 273 K	2.00 atm	500. mL	25.0°C 298 K
5	95.0 kPa	4.00 L		101. kPa	6.00 L	498.°C 471. K or
6	650. torr		100.°C 373 K	900. torr	225. mL	150.°C 423 K
7	850. mmHg	1.50 L	15.0°C 288 K		2.50 L	30.0°C 303 K
8	125. kPa	125. mL		100. kPa	100 mL	75.0°C 348 K

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CL Gary

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'CL Gary' text.

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