EXTRA PRACTICE PROBLEMS FOR MODULE #6

I. Nitric acid is manufactured with the following reaction:

How many grams of nitrogen dioxide are required to produce 5.89x103 kg HNO3 in excess water?

2. Tungsten metal, W, is produced by the following reaction:

$$WO_3 + 3H_2 \rightarrow W + 3H_2O$$

How many grams of tungsten can be made from 1.0x10° g of hydrogen in excess WO₃?

3. How many grams of WO₃ would be used in the above reaction under the conditions specified in 82?

4. Carbon tetrachloride is made according to the following reaction:

How many grams of CS2 are needed to make 16 kg of CCL?

To make a phosphorus fertilizer (CaH₄P₂O₈), agricultural companies use the following reaction:

$$Ca_3P_2O_8 + 2H_2SO_4 + 4H_2O \rightarrow CaH_4P_2O_8 + 2CaH_4SO_6$$

If 2.50x103 grams of Ca₃P₂O₈ are reacted with excess H₂SO₄ and H₂O, how many arms of fertilizer can be made?

6. In the following reaction:

$$Mg(s) + 2H_2O(l) \rightarrow Mg(OH)_2(s) + H_2(g)$$

What mass of Mg must you use to make 5.6x10° g of Mg(OH);?

7. Which of the following are empirical formulas?

a. SO:

b. K-S-O-

c. CsH₁₂Os

d. PaHi

8. Decomposition of 100.0 g of a dark green powder yields 39.7 g K, 27.9 g Mn, and an unknown amount of oxygen. What is the empirical formula?

9. The mass of the molecule above is 197.1 amu. What is the molecular formula?

10. A compound has an empirical formula of CH2O. If its molecular mass is 60.0 amu, what is its molecular formula?

Exploring Creation With Chemistry Practice Problem Answers

JS Bruner

Exploring Creation With Chemistry Practice Problem Answers:

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Table of Contents Exploring Creation With Chemistry Practice Problem Answers

- 1. Understanding the eBook Exploring Creation With Chemistry Practice Problem Answers
 - The Rise of Digital Reading Exploring Creation With Chemistry Practice Problem Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Exploring Creation With Chemistry Practice Problem Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Exploring Creation With Chemistry Practice Problem Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Exploring Creation With Chemistry Practice Problem Answers

- Personalized Recommendations
- Exploring Creation With Chemistry Practice Problem Answers User Reviews and Ratings
- Exploring Creation With Chemistry Practice Problem Answers and Bestseller Lists
- 5. Accessing Exploring Creation With Chemistry Practice Problem Answers Free and Paid eBooks
 - Exploring Creation With Chemistry Practice Problem Answers Public Domain eBooks
 - Exploring Creation With Chemistry Practice Problem Answers eBook Subscription Services
 - Exploring Creation With Chemistry Practice Problem Answers Budget-Friendly Options
- 6. Navigating Exploring Creation With Chemistry Practice Problem Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Exploring Creation With Chemistry Practice Problem Answers Compatibility with Devices
 - Exploring Creation With Chemistry Practice Problem Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Exploring Creation With Chemistry Practice Problem Answers
 - Highlighting and Note-Taking Exploring Creation With Chemistry Practice Problem Answers
 - Interactive Elements Exploring Creation With Chemistry Practice Problem Answers
- 8. Staying Engaged with Exploring Creation With Chemistry Practice Problem Answers
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Exploring Creation With Chemistry Practice Problem Answers
- 9. Balancing eBooks and Physical Books Exploring Creation With Chemistry Practice Problem Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Exploring Creation With Chemistry Practice Problem Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Exploring Creation With Chemistry Practice Problem Answers
 - Setting Reading Goals Exploring Creation With Chemistry Practice Problem Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Exploring Creation With Chemistry Practice Problem Answers

- Fact-Checking eBook Content of Exploring Creation With Chemistry Practice Problem Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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