

Activity B:**Methane reaction**Get the Gizmo ready:

- Select $C + H_2$ becomes CH_4 .



Goal: Predict the results of a chemical reaction.

1. Count: CH_4 is the chemical formula for methane. In order to produce a single molecule of methane, how many carbon (C) atoms are needed? Hydrogen atoms?

2. Predict: Set the number of C atoms to 14 and the number of H_2 molecules to 10.

A. How many total hydrogen atoms are there?

B. How many methane molecules could form from these reactants?

C. After the atoms react, which reactant will be left over?

D. Which reactant will be the limiting reactant?

E. Click **Play** and wait until **Reaction complete** is shown. What happened?

Each hydrogen molecule formed a connection with carbon (4 hydrogen molecules per carbon). There are extra carbon atoms that did not link because there are additional carbon atoms.

3. Test: Click **Reset**. Set the number of C atoms to seven.

A. How many hydrogen molecules (H_2) will be needed for there to be no excess reactants? Use the slider to set the number of H_2 molecules to this value.

B. How many methane molecules will be formed?

C. Click **Play**. What happened?

Since all of the hydrogen molecules linked to the carbon, there were no extra reactants.

4. Solve: For each **chemical equation** below, write the number of product molecules that will form from the reaction. Then, circle the limiting reactant. (Note: The **coefficients** in front of the reactants indicate the number of reactant molecules or atoms present.)



Explore Learning Limiting Reactants Actual Answer Key

RD Boyd



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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{Ag}^+ \text{aq} + 2\text{F}^- \text{aq} + 2\text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2\text{NH}_4\text{F} \text{aq}$ because two NH_4^+ aq

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