

Activity A:
Boyle's law

Get the Gizmo ready:

- Set the temperature (T) to 300 K.
- Check that the mass (m) is set to 0 kg.



Question: How does pressure affect the volume of a gas?

1. **Form hypothesis:** In this experiment, you will pile weights on the lid of the container of gas. What do you think will happen as more weight is added to the lid?

The harder it will be for the pressure to keep it up.

2. **Notice:** Look at the DESCRIPTION pane. What is the mass of the lid? **10 kg**

How much pressure does the lid exert on the gas? **98.1 N/m²**

3. **Collect data:** With the temperature held constant at 300 K, use the **Select mass** slider to place weights on the lid. Record the pressure and volume of the gas for each added mass.

Added mass on the lid	Total mass (lid + added mass)	Pressure*	Volume
0 kg	10 kg	98.1 N/m ²	2.54 m ³
10 kg	20 kg	196.2 N/m ²	1.27 m ³
20 kg	30 kg	294.3 N/m ²	0.85 m ³
30 kg	40 kg	392.4 N/m ²	0.64 m ³

*This model does not include atmospheric pressure, which is 101,325 N/m².

4. **Analyze:** As the pressure increases at constant temperature, what happens to the volume of the gas? **The volume lowers every time the pressure rises**

This relationship is called **Boyle's law**.

5. **Calculate:** Compare the pressure and volume values in your data table.

A. How did doubling the pressure change the gas volume? **It lowered the volume by half of the original**

B. How did tripling the pressure change the gas volume? **It lowered the original volume by 2/4 of its value**

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