

**Activity A:**  
**Boyle's law**

Get the Gizmo ready:

- Set the temperature ( $T$ ) to 300 K.
- Check that the mass ( $m$ ) is set to 0 kg.



**Question:** How does pressure affect the volume of a gas?

1. **Form hypothesis:** In this experiment, you will pile weights on the lid of the container of gas. What do you think will happen as more weight is added to the lid?

**The harder it will be for the pressure to keep it up.**

2. **Notice:** Look at the DESCRIPTION pane. What is the mass of the lid? **10 kg**

How much pressure does the lid exert on the gas? **98.1 N/m<sup>2</sup>**

3. **Collect data:** With the temperature held constant at 300 K, use the **Select mass** slider to place weights on the lid. Record the pressure and volume of the gas for each added mass.

| Added mass on the lid | Total mass (lid + added mass) | Pressure*              | Volume              |
|-----------------------|-------------------------------|------------------------|---------------------|
| 0 kg                  | 10 kg                         | 98.1 N/m <sup>2</sup>  | 2.54 m <sup>3</sup> |
| 10 kg                 | 20 kg                         | 196.2 N/m <sup>2</sup> | 1.27 m <sup>3</sup> |
| 20 kg                 | 30 kg                         | 294.3 N/m <sup>2</sup> | 0.85 m <sup>3</sup> |
| 30 kg                 | 40 kg                         | 392.4 N/m <sup>2</sup> | 0.64 m <sup>3</sup> |

\*This model does not include atmospheric pressure, which is 101,325 N/m<sup>2</sup>.

4. **Analyze:** As the pressure increases at constant temperature, what happens to the volume of the gas? **The volume lowers every time the pressure rises.**

This relationship is called **Boyle's law**.

5. **Calculate:** Compare the pressure and volume values in your data table.

A. How did doubling the pressure change the gas volume? **It lowered the volume by half of the original.**

B. How did tripling the pressure change the gas volume? **It lowered the original volume by 2/4 of its value.**

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