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web answer a 1 explanation the entire reaction was multiplied by 2 therefore  $\Delta H$  must also be halved 3 answer c explanation option c is incorrect because 566 kJ/mol is the  $\Delta H$  per mole of  $\text{O}_2$  and not for 2 moles even though options b and d are expressed in grams they are still correct because 56

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 $2S_3O_2 + 2SO_3 \rightarrow 3H_2O + 7914kJ$  is this endothermic or exothermic?

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 $2H_2(g) + O_2(g) \rightarrow 2H_2O(l)$  energy endothermic the transfer of heat from the surroundings to the system energy  
 $2Hg(s) + 2Hg(l) \rightarrow O_2(g)$

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web  $q = 140\text{ g} \times 4.184\text{ J}/\text{g} \cdot \text{C} = 585.76\text{ J}$   
 $q = 9372.16\text{ J}$   
 $35.5\text{ g} \times 1\text{ mole KCN} = 0.5453\text{ moles}$   
 $65.1\text{ g} = 0.5453\text{ mole moles}$   
 5 when a 15 g sample of  $SnCl_2$  dissolves in 80 g of water in a calorimeter the temperature drops from  $42^\circ\text{C}$  to  $15^\circ\text{C}$  calculate the  $\Delta h$  for the process  
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