

SOLUTIONS TO PROBLEMS

Chapter 1

- An amino or imino group gives a molecule a positive charge; a carboxylate or phosphoryl group gives a molecule a negative charge.
- A** Thiol (sulfhydryl) group
B Carbonyl group
C Amide linkage
D Phosphoanhydride (pyrophosphoryl) linkage
E Phosphoryl group (P_i)
F Hydroxyl group
- Hydrolysis reactions tend to occur with an increase in entropy, because the highly ordered polymer is broken down into separate units.
- Typical prokaryotic cells are 1–10 μm in diameter, so *T. acidophilus* is about 10–300 times larger (or has 1000 to 27 million times the volume). Typical eukaryotic cells are 10–100 μm in diameter, so *T. acidophilus* is about the same size or up to 30 times larger (or has the same to 27,000 times the volume).
- The data indicate that eukaryotes are more similar to the archaea than to the eubacteria.
- (a) Lipid water; (b) ice has less entropy at the lower temperature.
- (a) Decreases; (b) increases; (c) increases; (d) no change.
- No. When the change in enthalpy is positive and the change in entropy is negative, the free energy change for the process is greater than zero, which makes the process unfavorable.
- (a) False. A spontaneous reaction occurs in only one direction.
(b) False. Thermodynamics does not specify the rate of a reaction.
(c) True. (d) True. A reaction is spontaneous as long as $\Delta S > \Delta H/T$.
- (a) $T = 273 + 10 = 283 \text{ K}$
 $\Delta G = \Delta H - T\Delta S$
 $\Delta G = 15 \text{ kJ} - (283 \text{ K})(0.05 \text{ kJ} \cdot \text{K}^{-1})$
 $= 15 - 14.15 \text{ kJ} = 0.85 \text{ kJ}$
 ΔG is greater than zero, so the reaction is not spontaneous.
(b) $T = 273 + 80 = 353 \text{ K}$
 $\Delta G = \Delta H - T\Delta S$
 $\Delta G = 15 \text{ kJ} - (353 \text{ K})(0.05 \text{ kJ} \cdot \text{K}^{-1})$
 $= 15 - 17.65 \text{ kJ} = -2.65 \text{ kJ}$
 ΔG is less than zero, so the reaction is spontaneous.
- $\Delta G = \Delta H - T\Delta S$
 $\Delta G = -7000 \text{ J} \cdot \text{mol}^{-1} - (298 \text{ K})(-25 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1})$
 $\Delta G = -7000 + 7450 \text{ J} \cdot \text{mol}^{-1} = 450 \text{ J} \cdot \text{mol}^{-1}$
The reaction is not spontaneous because $\Delta G > 0$. The temperature must be decreased in order to decrease the value of the $T\Delta S$ term.
- $\Delta G^\circ = -RT \ln \frac{[C]}{[A][B]}$
 $= -(8.314 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1})(298 \text{ K}) \ln \frac{(0.009)}{(0.002)(0.003)}$
 $= -18,000 \text{ J} \cdot \text{mol}^{-1} = -18 \text{ kJ} \cdot \text{mol}^{-1}$
- $\Delta G^\circ = -RT \ln K_{eq} = -RT \ln \frac{[C][D]}{[A][B]}$
 $= -(8.314 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1})(298 \text{ K}) \ln \left(\frac{3 \times 10^{-3}}{(5 \times 10^{-2})^2} \right)$
 $= 5700 \text{ J} \cdot \text{mol}^{-1} = 5.7 \text{ kJ} \cdot \text{mol}^{-1}$

Since ΔG° is positive, the reaction is endergonic under standard conditions.

- $K_{eq} = e^{-\Delta G^\circ/RT} = e^{-(-20000 \text{ J} \cdot \text{mol}^{-1})/(8.314 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1})(298 \text{ K})}$
 $= 4.6 \times 10^3$
- From Eq. 1-17, $K_{eq} = \frac{[\text{G6P}][\text{G1P}]}{[\text{G6P}][\text{G1P}]} = e^{-\Delta G^\circ/RT}$
 $\frac{[\text{G6P}][\text{G1P}]}{[\text{G6P}][\text{G1P}]} = e^{-(-17000 \text{ J} \cdot \text{mol}^{-1})/(8.314 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1})(298 \text{ K})}$
 $\frac{[\text{G6P}][\text{G1P}]}{[\text{G1P}][\text{G6P}]} = 17.6$
- The cell membrane must be semipermeable so that the cell can retain essential compounds while allowing nutrients to enter and wastes to exit.
- Membrane-enclosed cellular compartments are typical of eukaryotic but not prokaryotic cells.
- Concentration = (number of moles)/(volume)
Volume = $(4/3)\pi r^3 = (4/3)\pi(5 \times 10^{-7} \text{ m})^3$
 $= 5.24 \times 10^{-22} \text{ m}^3 = 5.24 \times 10^{-16} \text{ L}$
Moles of protein = (2 molecules)/(5.24 × 10⁻¹⁶ L)
 $= 3.82 \times 10^{-16} \text{ mol}$
Concentration = $(3.82 \times 10^{-16} \text{ mol})/(5.24 \times 10^{-16} \text{ L})$
 $= 6.3 \times 10^{-1} \text{ M} = 6.3 \text{ mM}$
- Number of molecules = (molar conc.)(volume) $(6.022 \times 10^{23} \text{ molecules} \cdot \text{mol}^{-1})$
 $= (1.0 \times 10^{-3} \text{ mol} \cdot \text{L}^{-1})(5.24 \times 10^{-16} \text{ L})$
 $(6.022 \times 10^{23} \text{ molecules} \cdot \text{mol}^{-1})$
 $= 3.2 \times 10^7 \text{ molecules}$
- In order for ΔG to have a negative value (a spontaneous reaction), $T\Delta S$ must be greater than ΔH .
 $T\Delta S > \Delta H$
 $T > \Delta H/\Delta S$
 $T > 7000 \text{ J} \cdot \text{mol}^{-1}/(20 \text{ J} \cdot \text{K}^{-1} \cdot \text{mol}^{-1})$
 $T > 350 \text{ K}$ or 77°C
- (a) Because $K_{eq} = \frac{[R]}{[Q]} = 25$, at equilibrium, the concentration of R is 25 times greater than the concentration of Q. When equal concentrations of Q and R are mixed, molecules of Q will be converted to molecules of R.
(b) Let x = amount of Q converted to R, so that [R] will be $50 \mu\text{M} + x$ and [Q] will be $50 \mu\text{M} - x$. Since $K_{eq} = \frac{[R]}{[Q]} = 25$,
$$\frac{50 + x}{50 - x} = 25(50 - x)$$

$$50 + x = 1250 - 25x$$

$$26x = 1200$$

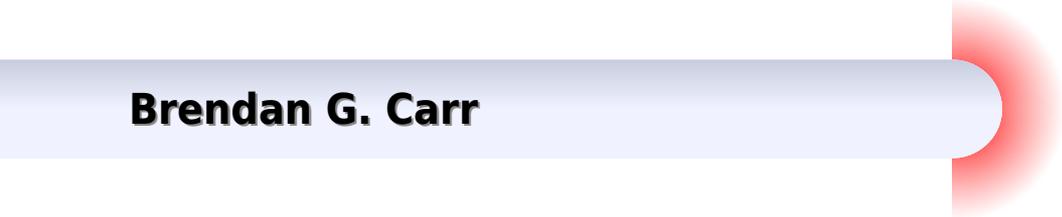
$$x = 46.15$$

$$[R] = 50 \mu\text{M} + 46.15 \mu\text{M} = 96.15 \mu\text{M}$$

$$[Q] = 50 \mu\text{M} - 46.15 \mu\text{M} = 3.85 \mu\text{M}$$
- At $10^\circ\text{C} = 283 \text{ K}$ ($1/T = 0.00353$), $K_{eq} = 100$ and $\ln K = 4.61$.
At $30^\circ\text{C} = 303 \text{ K}$ ($1/T = 0.00330$), $K_{eq} = 10$ and $\ln K = 2.30$.
These two points generate a line on a van't Hoff plot (ln K_{eq} versus $1/T$) with a positive slope that is equal to $(-\Delta H^\circ/R)$. ΔH° must therefore be negative, indicating that enthalpy decreases during the reaction (heat is given off).

Fundamentals Of Biochemistry Voet And Solution Manual

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