

Name : _____ Date : _____

Score : _____



Electron Configuration



1. Determine what elements are denoted by the following configurations.

i. $1s^2 2s^2 2p^6 3s^2 3p^4$ _____

ii. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$ _____

iii. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$ _____

iv. $[\text{Kr}] 5s^2 4d^{10} 5p^3$ _____

v. $[\text{Ne}] 3s^2 3p^1$ _____

2. Write the full electron configuration (ex., $1s^2 2s^1$) of the following elements?

i. Nitrogen _____

ii. Silicon _____

iii. Calcium _____

iv. Strontium _____

v. Bismuth _____

3. Write the abbreviated electron configuration (ex., $[\text{He}] 2s^1$) of the following element?

i. Fluorine _____

ii. Magnesium _____

iii. Selenium _____

iv. Rubidium _____

v. Lead _____

4. If each orbital can hold a maximum of two electrons, how many can each of the following hold?

i. $2s$ _____ ii. $5p$ _____ iii. $4f$ _____ iv. $3d$ _____ v. $4d$ _____

5. What is the shape of an s orbital? _____

6. What is the shape of a p orbital? _____

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BM King



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Table of Contents Matter And The Periodic Table Writing Electron Configurations Worksheet

1. Understanding the eBook Matter And The Periodic Table Writing Electron Configurations Worksheet
 - The Rise of Digital Reading Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Matter And The Periodic Table Writing Electron Configurations Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Personalized Recommendations
 - Matter And The Periodic Table Writing Electron Configurations Worksheet User Reviews and Ratings
 - Matter And The Periodic Table Writing Electron Configurations Worksheet and Bestseller Lists
5. Accessing Matter And The Periodic Table Writing Electron Configurations Worksheet Free and Paid eBooks
 - Matter And The Periodic Table Writing Electron Configurations Worksheet Public Domain eBooks
 - Matter And The Periodic Table Writing Electron Configurations Worksheet eBook Subscription Services
 - Matter And The Periodic Table Writing Electron Configurations Worksheet Budget-Friendly Options
6. Navigating Matter And The Periodic Table Writing Electron Configurations Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Matter And The Periodic Table Writing Electron Configurations Worksheet Compatibility with Devices
 - Matter And The Periodic Table Writing Electron Configurations Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Highlighting and Note-Taking Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Interactive Elements Matter And The Periodic Table Writing Electron Configurations Worksheet

8. Staying Engaged with Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Matter And The Periodic Table Writing Electron Configurations Worksheet
9. Balancing eBooks and Physical Books Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Matter And The Periodic Table Writing Electron Configurations Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Setting Reading Goals Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Fact-Checking eBook Content of Matter And The Periodic Table Writing Electron Configurations Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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