

- Instructions: (1) Substantiate your answers with formulas for finding using one or more of the following methods:
 (2) Balance each equation.
 (3) Identify the type reaction: synthesis, decomposition, single replacement, double replacement, or combustion.

1. Aluminum metal reacts with oxygen (air) to form aluminum oxide and oxygen diatomic molecules.

Type: single replacement

Reaction: $4\text{Al(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Al}_2\text{O}_3\text{(s)}$

2. Iron(III) sulfate solution and barium ions (Ba) sulfate.

Type: double replacement

Reaction: $\text{Fe}_2(\text{SO}_4)_3 + 3\text{Ba}^{2+} \rightarrow 2\text{Fe}^{3+} + 3\text{BaSO}_4$

3. Calciumphosphorus pentoxide can be produced from the elements oxygen and phosphorus. Write down chemical phosphorus reactions that are part of the reaction.

Type: double replacement

Reaction: $\text{P}_4 + 5\text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$

4. Iron(III)phosphate reacts to produce iron(III) sulfate and phosphoric acid.

Type: double replacement

Reaction: $\text{Fe}_3(\text{PO}_4)_2 + 4\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 2\text{H}_3\text{PO}_4$

5. Iron(III)phosphate sulfate reacts to produce iron(III) sulfate, iron(III) phosphate, and sulfur dioxide.

Type: double replacement

Reaction: $\text{Fe}_3(\text{PO}_4)_2 + 4\text{H}_2\text{SO}_4 \rightarrow 2\text{Fe}_2(\text{SO}_4)_3 + 2\text{H}_3\text{PO}_4$

Even More Fun With Equations 6 33

Katrin Zwirgmaier



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