

Unit 18: Hydrocarbon and its Compounds

Things to be Studied in this Unit

- Introduction to hydrocarbon
- Saturated and unsaturated hydrocarbon
- Alkane, alkene and alkyne
- Homologous series
- Alkyl radical
- Functional group
- IUPAC system
- Common system to nomenclature of hydrocarbons
- Some hydrocarbon and its compounds
 - Methane
 - Propane
 - Alcohol
 - Ethane
 - Butane
 - Glycerol

Introduction

Hydrocarbons are organic compounds composed of hydrogen (H) and carbon (C) atoms. They are of two main types, viz. saturated and Unsaturated. The hydrocarbons in which carbon atoms are combined with a single bond are called saturated hydrocarbons. The hydrocarbons which contain a double bond or triple bond between carbon atoms are called unsaturated hydrocarbons. Hydrocarbons are the basis for many organic compounds and are found in various forms in fossil fuels like mineral oil and natural gas. They serve as essential building blocks for a wide range of chemical products and fuels, playing a vital role in industry and energy production.

Things to Know

-  **Hydrocarbon:** Hydrocarbons are the compounds formed by the combination of hydrogen and carbon. Examples: Methane (CH_4), Propane (C_3H_8) etc.
-  **Unsaturated hydrocarbon:** The hydrocarbon which contains a double bond or triple bond between carbon atoms is called unsaturated hydrocarbon. The general formula of unsaturated hydrocarbons is C_nH_{2n} (alkene) and $\text{C}_n\text{H}_{2n-2}$ (alkyne).
-  **Saturated hydrocarbon:** The hydrocarbon in which carbon atoms are combined with a single bond is called saturated hydrocarbon. The general formula of saturated hydrocarbon is $\text{C}_n\text{H}_{2n+2}$.
-  **IUPAC:** IUPAC stands for International Union of Pure and Applied Chemistry.
-  **Alkane:** An alkane is a saturated hydrocarbon in which carbon atoms are connected by only single covalent bonds. The general formula of alkanes is $\text{C}_n\text{H}_{2n+2}$ where 'n' is the number of carbon atoms present in one molecule of alkane. Example: Propane, Pentane, etc.
-  **Alkene:** An alkene is a saturated hydrocarbon in which carbon atoms are connected by a double bond. So an alkene contains $-\overset{\textstyle |}{\underset{\textstyle |}{\text{C}}}=\overset{\textstyle |}{\underset{\textstyle |}{\text{C}}}-$ group. The general formula of an alkene is C_nH_{2n} where 'n' is the number of carbon atoms in its one molecule. Example: Propene, Pentene, etc.

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Table of Contents Hydrocarbon Compounds Guide Answers

1. Understanding the eBook Hydrocarbon Compounds Guide Answers
 - The Rise of Digital Reading Hydrocarbon Compounds Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Hydrocarbon Compounds Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hydrocarbon Compounds Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hydrocarbon Compounds Guide Answers

- Personalized Recommendations
- Hydrocarbon Compounds Guide Answers User Reviews and Ratings
- Hydrocarbon Compounds Guide Answers and Bestseller Lists
- 5. Accessing Hydrocarbon Compounds Guide Answers Free and Paid eBooks
 - Hydrocarbon Compounds Guide Answers Public Domain eBooks
 - Hydrocarbon Compounds Guide Answers eBook Subscription Services
 - Hydrocarbon Compounds Guide Answers Budget-Friendly Options
- 6. Navigating Hydrocarbon Compounds Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Hydrocarbon Compounds Guide Answers Compatibility with Devices
 - Hydrocarbon Compounds Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hydrocarbon Compounds Guide Answers
 - Highlighting and Note-Taking Hydrocarbon Compounds Guide Answers
 - Interactive Elements Hydrocarbon Compounds Guide Answers
- 8. Staying Engaged with Hydrocarbon Compounds Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hydrocarbon Compounds Guide Answers
- 9. Balancing eBooks and Physical Books Hydrocarbon Compounds Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hydrocarbon Compounds Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hydrocarbon Compounds Guide Answers
 - Setting Reading Goals Hydrocarbon Compounds Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hydrocarbon Compounds Guide Answers

- Fact-Checking eBook Content of Hydrocarbon Compounds Guide Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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web play a game to revise and reinforce prior learning on magnetic forces explore how magnets behave towards one another in a wide variety of different situations form