

Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reagent

$$14.8 \text{ g C}_3\text{H}_8 \times \frac{1 \text{ mol C}_3\text{H}_8}{44.09 \text{ g C}_3\text{H}_8} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.2 \text{ g H}_2\text{O}$$

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 1.55 \text{ g H}_2\text{O}$$

O_2 is limiting

b) determine the number of moles of carbon dioxide produced

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol CO}_2} = 2.84 \text{ g H}_2\text{O}$$

c) determine the number of grams of H_2O produced

$$1.55 \text{ g H}_2\text{O}$$

d) determine the number of grams of excess reagent left

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{1 \text{ mol C}_3\text{H}_8}{5 \text{ mol O}_2} \times \frac{44.09 \text{ g C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} = .948 \text{ g C}_3\text{H}_8 \text{ used}$$

$$14.8 \text{ g C}_3\text{H}_8 - 0.948 \text{ g C}_3\text{H}_8 = 13.9 \text{ g C}_3\text{H}_8 \text{ remain}$$

2. Given the following equation:



a) If 10.0 g of $\text{Al}_2(\text{SO}_4)_3$ is reacted with 10.0 g of NaOH , determine the limiting reagent

$$10.00 \text{ g Al}_2(\text{SO}_4)_3 \times \frac{1 \text{ mol Al}_2(\text{SO}_4)_3}{294.17 \text{ g Al}_2(\text{SO}_4)_3} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 12.9 \text{ g Na}_2\text{SO}_4$$

$$10.00 \text{ g NaOH} \times \frac{1 \text{ mol NaOH}}{24.17 \text{ g NaOH}} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{6 \text{ mol NaOH}} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 15.8 \text{ g Na}_2\text{SO}_4$$

Limiting Reagent 2 Key

Christian Drosten

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Limiting Reagent 2 Key

1. Understanding the eBook Limiting Reagent 2 Key
 - The Rise of Digital Reading Limiting Reagent 2 Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent 2 Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagent 2 Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent 2 Key
 - Personalized Recommendations
 - Limiting Reagent 2 Key User Reviews and Ratings
 - Limiting Reagent 2 Key and Bestseller Lists

5. Accessing Limiting Reagent 2 Key Free and Paid eBooks
 - Limiting Reagent 2 Key Public Domain eBooks
 - Limiting Reagent 2 Key eBook Subscription Services
 - Limiting Reagent 2 Key Budget-Friendly Options
6. Navigating Limiting Reagent 2 Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reagent 2 Key Compatibility with Devices
 - Limiting Reagent 2 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reagent 2 Key
 - Highlighting and Note-Taking Limiting Reagent 2 Key
 - Interactive Elements Limiting Reagent 2 Key
8. Staying Engaged with Limiting Reagent 2 Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reagent 2 Key
9. Balancing eBooks and Physical Books Limiting Reagent 2 Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagent 2 Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reagent 2 Key
 - Setting Reading Goals Limiting Reagent 2 Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reagent 2 Key
 - Fact-Checking eBook Content of Limiting Reagent 2 Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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